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web where the salt is the conjugate base of the acid the four parts of the titration curve are described below and you should look to the appropriate text section to see how they are treated pure acid 0 ml of base is added section 17.3.2.1

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web solution a titrant volume 0 ml the solution ph is due to the acid ionization of HCl because this is a strong acid the ionization is complete and the hydronium ion molarity is 0.100 M the ph of the solution is then $-\log 0.100 = 1.000$ ph $-\log 0.100 = 1.000$ b titrant volume 12.50 ml

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web expert answer 100% 2 ratings the example is based upon henderson hasselbalch equation the equation gives relation between ph and pK_a of weak acid the equation is $ph = pK_a + \log \frac{[salt]}{[acid]}$ at equivalence point all a view the full answer transcribed image text

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web ab acid base titration curves purpose the purpose of this lab is to determine the shape of acid base titration curves theoretically and experimentally and to use an acid base titration to measure the effectiveness of antacid tablets summary a titration curve is the plot of the pH value of a solution vs the amount of titrant added

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web pH titration graph 1 lab data provide columns a and b 2 a graph of pH column b plotted as ordinate vs volume column a as abscissa provides a titration curve as illustrated by graph 1 or 2 below 3 estimate of the midpoint of the steeply rising section of the titration curve provides an evaluation of the equivalence point

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web acid base titration lab report jessica chung chem 267 lab date completed 25 september 2019 date turned in 9 october 2019 figure 1 acid base titration curve from titration with indicator for part 1 volume of NaOH added mL pH of the solution

Volume of NaOH added (mL)	pH
4	6.4
6	4.6
5	9.5
9	5.9

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web background plot of the pH of a solution against the volume of titrant added is called a titration curve the pH can be measured directly with a pH meter while titrant is added from a buret for the acids used in this experiment the titrant will

always be a 0.1 M solution of the strong base NaOH

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web strong acids and strong bases completely ionize in solution resulting in water and a salt for a weak acid there is only partial ionization the titration curve of a strong base strong acid showed a stretched out curve as it started with a slow gradual change in pH as it reached the equivalence point