

CHAPTER 6 REVIEW*Chemical Bonding***SECTION 6-2**

SHORT ANSWER Answer the following questions in the space provided.

1. Use the concept of potential energy to describe how a covalent bond forms between two atoms.

2. Name two elements that form compounds that are exceptions to the octet rule.

3. Explain why resonance structures illustrate the limitations of Lewis structures in correctly modeling covalent bonds.

4. Bond energy is related to bond length. Use the data in the tables below to arrange the bonds listed in order of increasing bond length, from shortest bond to longest.

a.

Bond	Bond energy (kJ/mol)
H—F	569
H—I	299
H—Cl	432
H—Br	366

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