

## Chemistry

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1A

## FUNDAMENTAL CONCEPTS

5

2011

- Q.1 In mass spectrometer, detector or collector measures the:  
 A) Masses of isotopes  
 B) Percentages of isotopes  
 C) Relative abundances of isotopes  
 D) Mass numbers of isotopes
- Q.2 How many  $^{35}\text{Cl}$  (chlorine) atoms are in two moles of chlorine?  
 A)  $2 \times 6.02 \times 10^{23}$  atoms  
 B)  $35.5 \times 6.02 \times 10^{23}$  atoms  
 C)  $2 \times 16^{23}$  atoms  
 D)  $2 \times 6.02 \times 10^{23}$  atoms

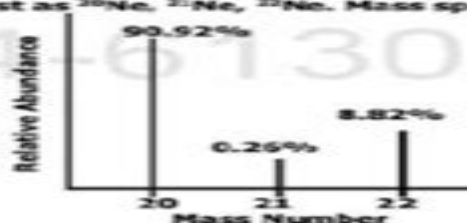
2012

- Q.3 An organic compound has empirical formula  $\text{C}_3\text{H}_5\text{O}$ , if molar mass of compound is  $110.15 \text{ g mol}^{-1}$ . The molecular formula of this organic compound is (A, of C=12, H=1.008 and O=16)  
 A)  $\text{C}_6\text{H}_{10}\text{O}_2$   
 B)  $\text{C}_9\text{H}_{15}\text{O}$   
 C)  $\text{C}_6\text{H}_8\text{O}_2$   
 D)  $\text{C}_9\text{H}_{15}\text{O}_2$
- Q.4 When 8 grams (4 moles) of  $\text{H}_2$  react with 2 moles of  $\text{O}_2$ , how many moles of water will be formed?  
 A) Five  
 B) Four  
 C) Six  
 D) Three

2013

- Q.5 Hydrogen burns in chlorine to produce hydrogen chloride. The ratio of masses of reactants in chemical reaction is:  

$$\text{H}_2 + \text{Cl}_2 \rightarrow 2\text{HCl}$$
  
 A) 1:35.5  
 B) 2:35.5  
 C) 1:71  
 D) 2:70
- Q.6 A sample of Neon is found to exist as  $^{20}\text{Ne}$ ,  $^{21}\text{Ne}$ ,  $^{22}\text{Ne}$ . Mass spectrum of 'Ne' is as follow:

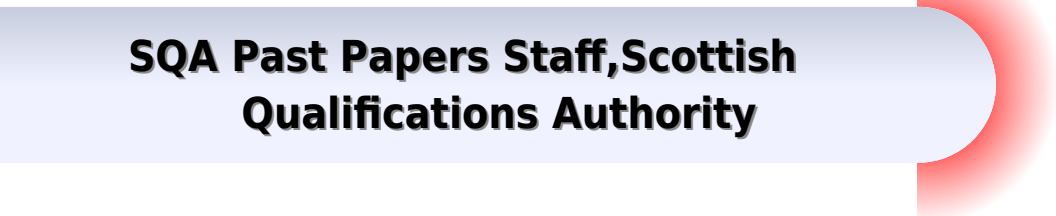


What is the relative atomic mass (A, value) of Neon?

- A) 20.18  
 B) 20.28  
 C) 20.10  
 D) 20.22

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