

2.2 Properties of Water

Lesson Objectives

-  Discuss the unique properties of water.
-  Differentiate between solutions and suspensions.
-  Explain what acidic solutions and basic solutions are.

Lesson Summary

The Water Molecule Water molecules (H_2O) are polar because of an uneven distribution of electrons, creating a slight negative (–) charge in the oxygen atom and a slight positive (+) charge in each hydrogen atom. The attraction between a hydrogen atom of one water molecule and the oxygen atom of another water molecule is called a **hydrogen bond**.

- ▶ **Cohesion** is an attraction between molecules of the same substance. It causes water molecules to be drawn together, producing surface tension.
- ▶ **Adhesion** is an attraction between molecules of different substances. It causes capillary action, an effect that causes water to rise in a narrow tube against the force of gravity.

Solutions and Suspensions A **mixture** is a material composed of two or more elements or compounds that are physically mixed together but not chemically combined. A **solution** is a mixture in which all the components are evenly spread out: the substance dissolved is the **solute**; the substance that causes the dissolving is the **solvent**. Mixtures of water and undissolved materials are **suspensions**.

Acids, Bases, and pH A water molecule (H_2O) can split apart to form a hydrogen ion (H^+) and a hydroxide ion (OH^-).

- ▶ The **pH scale** measures the concentration of hydrogen ions in a solution. The scale ranges from 0 to 14. Pure water has a pH of 7.
- ▶ An **acid** is any compound that forms H^+ ions in solution. Acidic solutions have pH values below 7. A **base** is a compound that forms OH^- ions in solution. Basic, or alkaline, solutions have pH values above 7.
- ▶ **Buffers** are weak acids or bases that can react with strong acids or bases to prevent sudden changes in pH.

The Water Molecule

For Questions 1–4, write True or False on the line provided.

- _____ 1. Water is a polar molecule.
- _____ 2. Hydrogen bonds are an example of adhesion.
- _____ 3. Covalent bonds give water a low heat capacity.
- _____ 4. A hydrogen bond is stronger than a covalent bond.

2 2 Properties Of Water Answers

Yijin Wang



2 2 Properties Of Water Answers:

The book delves into 2 2 Properties Of Water Answers. 2 2 Properties Of Water Answers is a vital topic that needs to be grasped by everyone, from students and scholars to the general public. This book will furnish comprehensive and in-depth insights into 2 2 Properties Of Water Answers, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:

- Chapter 1: Introduction to 2 2 Properties Of Water Answers
- Chapter 2: Essential Elements of 2 2 Properties Of Water Answers
- Chapter 3: 2 2 Properties Of Water Answers in Everyday Life
- Chapter 4: 2 2 Properties Of Water Answers in Specific Contexts
- Chapter 5: Conclusion

2. In chapter 1, this book will provide an overview of 2 2 Properties Of Water Answers. This chapter will explore what 2 2 Properties Of Water Answers is, why 2 2 Properties Of Water Answers is vital, and how to effectively learn about 2 2 Properties Of Water Answers.
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6. In chapter 5, this book will draw a conclusion about 2 2 Properties Of Water Answers. This chapter will summarize the key points that have been discussed throughout the book.

This book is crafted in an easy-to-understand language and is complemented by engaging illustrations. It is highly recommended for anyone seeking to gain a comprehensive understanding of 2 2 Properties Of Water Answers.

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