

9) A 250mL solution of 0.5M NaF has 50mL of a 1M BaF₂ solution added to it. What is the concentration of [F⁻] in the new solution? Assume that volume is additive.

From (8)

$$[F^-] = 0.125 \text{ mol} \left| \begin{array}{l} 0.05 \text{ L} \times 1 \text{ M BaF}_2 \\ = 0.05 \text{ mol BaF}_2 \\ = 0.1 \text{ mol F}^- \end{array} \right| \begin{array}{l} \text{Final sol'n} \\ \text{mol [F]}^- = 0.125 \text{ mol} + 0.1 \text{ mol} = 0.225 \text{ mol F}^- \\ \text{volume} = 0.25 \text{ L} + 0.05 \text{ L} \\ = 0.3 \text{ L} \end{array}$$

$$M = \frac{0.225 \text{ mol}}{0.3 \text{ L}} = \boxed{0.75 \text{ M}}$$

10) How many grams of NaHCO₃ (molar mass = 84.01 g/mol) are needed to make a 350mL solution of 0.75M NaHCO₃?

$$n = V \times M$$

$$= 0.35 \text{ L} \times 0.75 \text{ M NaHCO}_3$$

$$= 0.2625 \text{ mol NaHCO}_3$$

$$\text{mass} = 0.2625 \text{ mol NaHCO}_3 \times \frac{84.01 \text{ g NaHCO}_3}{1 \text{ mol NaHCO}_3}$$

$$= \boxed{22.05 \text{ g NaHCO}_3}$$

11) Calculate the amount of distilled water that is needed to dilute a 3L solution of 5M HBr to a concentration of 1.5M HBr?

$$\text{mol} = 3 \text{ L} \times 5 \text{ M}$$

$$= 15 \text{ mol HBr}$$

$$\text{volume needed: } V = \frac{15 \text{ mol}}{1.5 \text{ M HBr}}$$

$$= 10 \text{ L}$$

$$10 \text{ L} - 3 \text{ L} = \boxed{7 \text{ L needed to add}}$$

12) What volume of 8M H₂SO₄ is needed to prepare a 100mL solution of 0.5M H₂SO₄?

$$0.1 \text{ L} \times 0.5 \text{ M H}_2\text{SO}_4$$

$$= 0.005 \text{ mol H}_2\text{SO}_4 \text{ needed}$$

$$V_{\text{ol}} = \frac{0.005 \text{ mol H}_2\text{SO}_4}{8 \text{ M H}_2\text{SO}_4}$$

$$= 6.25 \times 10^{-4} \text{ L}$$

$$= \boxed{0.625 \text{ mL of } 8 \text{ M H}_2\text{SO}_4, \text{ the rest is water}}$$

13) The maximum amount of salt that can be contained in water is approximately 26.3% by weight. Convert this into concentration with units of molarity. Info: 1g of water = 1 mL of water.

$$\text{assume } 100 \text{ g}$$

$$26.3 \text{ g NaCl}$$

$$26.3 \text{ g NaCl} \times \frac{1 \text{ mol NaCl}}{58.44 \text{ g NaCl}} = 0.45 \text{ mol NaCl}$$

$$73.7 \text{ g H}_2\text{O} \times \frac{1 \text{ mol H}_2\text{O}}{18 \text{ g H}_2\text{O}} = 4.1 \text{ mol H}_2\text{O}$$

$$M = \frac{0.45 \text{ mol NaCl}}{0.0737 \text{ L sol'n}} = \boxed{6.1 \text{ M}}$$

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Richard Samuel Moog



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