

Chemistry Study Guide

Periodic Table of the Elements

1 H 1.008 3 Li 6.941 4 Be 9.012 5 B 10.81 6 C 12.01 7 N 14.01 8 O 16.00 9 F 18.998 10 Ne 20.18
11 Na 22.99 12 Mg 24.31 13 Al 26.98 14 Si 28.09 15 P 30.97 16 S 32.07 17 Cl 35.45 18 Ar 39.95
19 K 39.10 20 Ca 40.08 21 Sc 44.96 22 Ti 47.88 23 V 50.94 24 Cr 52.00 25 Mn 54.94 26 Fe 55.85 27 Co 58.93 28 Ni 58.71 29 Cu 63.55 30 Zn 65.38 31 Ga 69.72 32 Ge 72.64 33 As 74.92 34 Se 78.96 35 Br 79.90 36 Kr 83.80
37 Rb 85.47 38 Sr 87.62 39 Y 88.91 40 Zr 91.22 41 Nb 92.91 42 Mo 95.94 43 Tc 98.91 44 Ru 101.07 45 Rh 102.91 46 Pd 106.42 47 Ag 107.87 48 Cd 112.41 49 In 114.82 50 Sn 118.71 51 Sb 121.76 52 Te 127.60 53 I 126.91 54 Xe 131.29
55 Cs 132.91 56 Ba 137.33 57 La 138.91 58 Ce 140.12 59 Pr 140.91 60 Nd 144.24 61 Pm 144.91 62 Sm 150.36 63 Eu 151.96 64 Gd 157.25 65 Tb 158.93 66 Dy 162.50 67 Ho 164.93 68 Er 167.26 69 Tm 168.93 70 Yb 173.05 71 Lu 174.97
72 Fr 223.02 73 Ra 226.03 74 Ac 227.03 75 Th 232.04 76 Pa 231.04 77 U 238.03 78 Np 237.05 79 Pu 244.06 80 Am 243.06 81 Cm 247.07 82 Bk 247.07 83 Cf 251.08 84 Es 252.08 85 Fm 257.09 86 Md 258.10 87 No 259.10 88 Lr 262.11
89 Uuo 289.10

Objective:

To understand the *structure of an atom*, and the *nature + properties* of subatomic particles ✎

Matter:

Anything that has **mass** and **occupies space** ✎

Examples: air, gases, liquids, solids!

Atoms:

Atoms are the **smallest, indivisible particle** ✎

- **Outer part** → *shell/orbit/energy level* → carries the electrons (e^-)
- **Inner part** → *nucleus* → houses the protons (p^+) and neutrons (n)
- Outer & Inner part → together these are the **subatomic particles**

Subatomic Particle Table:

Particle	Symbol	Mass (amu)	Charge	Location
Proton	p^+	1	+1	Nucleus
Neutron	n	1	0	Nucleus
Electron	e^-	~ 0	-1	Orbit/Shell

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RD Boyd



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