

Name _____ Section _____

Experiment 2**Advance Study Assignment: Properties of Systems in Chemical Equilibrium**

1. Methyl orange, HMO, is a common acid-base indicator. In solution it ionizes according to the equation:



If methyl orange is added to distilled water, the solution turns yellow. If a drop or two of 6 M HCl is added to the yellow solution, it turns red. If to that solution one adds a few drops of 6 M NaOH the color reverts to yellow.

- Why does adding 6 M HCl to the yellow solution of methyl orange tend to cause the color to change to red? (Note that in solution HCl exists as $\text{H}^{\text{+}}$ and $\text{Cl}^{\text{-}}$ ions.)
- Why does adding 6 M NaOH to the red solution tend to make it turn back to yellow? (Note that in solution NaOH exists as $\text{Na}^{\text{+}}$ and $\text{OH}^{\text{-}}$ ions. How does increasing $[\text{OH}^{\text{-}}]$ shift Reaction 3 in the discussion section? How would the resulting change in $[\text{H}^{\text{+}}]$ affect the dissociation reaction of HMO?)

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Biren Prasad



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