Name: Key

For each reaction:

- \* Assign Oxidation Numbers
- \* Write an Oxidation and Reduction Half Reaction
- \*Circle the Oxidizing Agent and Underline the Reducing Agent.

1) Fe<sup>-3</sup> + 5n → Fe<sup>-2</sup> + 5n<sup>-2</sup>

Oxidation: Snº -> Snº + 2e

Reduction: (Fe') . Is' -> Fe'

2) AlF, \* Bo -> Al \* BoF.

Oxidation: Bo? -> Bort - 2e

Reduction: (Al-3) + 3e - Al-

3) H<sub>2</sub>O<sub>2</sub> + H<sub>2</sub>SO<sub>2</sub> → H<sub>2</sub>O + H<sub>2</sub>SO<sub>2</sub>

Oxidation: 5-1 + 5-1 + 2e

Reduction: (O1) + Ie: + O-

4)  $H_1SO_1 + KMnO_4 \rightarrow MnSO_4 + K_1SO_4 + H_2O_5$ 

Oxidation: 5" + 5" + 2e

Reduction: (Mn-) + 5e - Mn-2

5) K2Cr2O7 + H2C2O4 + H2SO4 + K2SO4 + Cr2(SO4)3 + CO2 + H2O

Oxidation: C1 -> C-1 + 1e

Reduction: (Cris) . 3c -> Cris

# **Analyzing Oxidation Reduction Reactions Answers Sheet**

United States. National Bureau of Standards

#### **Analyzing Oxidation Reduction Reactions Answers Sheet:**

CliffsNotes AP Chemistry Bobrow Test Preparation Services,2009-02-09 The book itself contains chapter length subject reviews on every subject tested on the AP Chemistry exam as well as both sample multiple choice and free response questions at each chapter s end Two full length practice tests with detailed answer explanations are included in the book

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