

Name \_\_\_\_\_ answer key \_\_\_\_\_ Date \_\_\_\_\_

# Stoichiometry: Mass – Mole Worksheet

Based on the following chemical equation, please answer the following questions!



Calculate the amount of moles of NaCl needed to produce 75 g of FeCl<sub>3</sub>.

$$\frac{9 \text{ FeCl}_3}{1 \text{ mole FeCl}_3} \times \frac{1 \text{ mole FeCl}_3}{9 \text{ FeCl}_3} \times \frac{\text{mole NaCl}}{\text{mole FeCl}_3}$$

Calculate the amount of Fe<sub>2</sub> (SO<sub>4</sub>)<sub>3</sub> in grams produce 215 g Na<sub>2</sub>SO<sub>4</sub>.

$$\frac{9 \text{ Na}_2\text{SO}_4}{1 \text{ mole Na}_2\text{SO}_4} \times \frac{1 \text{ mole Na}_2\text{SO}_4}{9 \text{ Na}_2\text{SO}_4} \times \frac{\text{mole Fe}_2(\text{SO}_4)_3}{\text{mole Na}_2\text{SO}_4} \times \frac{9 \text{ Fe}_2(\text{SO}_4)_3}{1 \text{ mole Fe}_2(\text{SO}_4)_3}$$

Calculate the amount in grams of FeCl<sub>3</sub> produced with 95g of Fe (SO<sub>4</sub>)<sub>3</sub>.

$$\frac{9 \text{ Fe}_2(\text{SO}_4)_3}{1 \text{ mole Fe}_2(\text{SO}_4)_3} \times \frac{1 \text{ mole Fe}_2(\text{SO}_4)_3}{9 \text{ Fe}_2(\text{SO}_4)_3} \times \frac{\text{mole FeCl}_3}{\text{mole Fe}_2(\text{SO}_4)_3} \times \frac{9 \text{ FeCl}_3}{1 \text{ mole FeCl}_3}$$

Calculate the amount of moles of Fe<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub> needed to react with 123.5 g of NaCl

$$\frac{9 \text{ NaCl}}{1 \text{ mole NaCl}} \times \frac{1 \text{ mole NaCl}}{9 \text{ NaCl}} \times \frac{\text{mole Fe}_2(\text{SO}_4)_3}{\text{mole NaCl}}$$

Based on the following chemical equation, please answer the following questions!



$$\frac{9 \text{ Ca(HCO}_3)_2}{1 \text{ mole Ca(HCO}_3)_2} \times \frac{1 \text{ mole Ca(HCO}_3)_2}{9 \text{ Ca(HCO}_3)_2} \times \frac{\text{mole CaCO}_3}{\text{mole Ca(HCO}_3)_2}$$

# Answers For Study Guide Mass And Moles

**Leo J. Malone, Theodore O. Dolter**



## Answers For Study Guide Mass And Moles:

*Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual* Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications emphasizing the relevance of the material they are learning. This edition continues the end of chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter.

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**Study Guide and Full Solutions Manual** Susan McMurry, John McMurry, 2002-08 Contains a brief overview of every chapter, review of skills, self-tests and the answers and detailed solutions to all end-of-chapter problems in the textbook.

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