2 Atoms, ions, and compounds Answers to practice questions

Question	Answ	Answer				I	Marks	Guidance	
1 (a)	C							B1	
1 (b)	В							81	
1 (c)	D	D						81	
2 (a)	38							B1	
2 (b)	32							B1	
2 (c)	5							B1	
3 (a) (i)	23							B1	
3 (a) (ii)	⁵⁸ 26Fe							82	Time to the second seco
3 (b)	²⁴ 11Na							B2	
3 (c)	*** _{Ja} Sr							B2	
4 (a)	Isotopes are atoms of the same element with different numbers of neutrons and different masses.						ent	B1	
4 (b)	Relative isotopic mass is the mass of an isotope relative to one-twelfth of the mass of an atom of carbon-12.						B2		
4 (c)	71	71						B1	
4 (d)	protons		neur	trons ele	etrons	charge	e	B2	
	770	8		9	10	2-			
	Z ² AJ	AJ 13 14		4	10				
5 (a)	Relative atomic mass is the weighted mean mass of an atom of an element relative to one-twelfth of the mass of an atom of carbon-12.							B3	
5 (b)	$A_{i} = 32.$	09			10.00 Ca	en es son		B2	
5 (c)	protons			neutrons	elect	electrons		B1	
	275	16	16 18		11	5			
	"S" 16			16	16 18				
6 (a)	$3Mg(s) + N_2(g) \rightarrow Mg_3N_2(s)$							B1 × 2	
	1 mark for correct species and state symbols 1 mark for balance								l.
6 (b)	Ca(s) + 2H₂O(l) → Ca(OH)₂(aq) + H₂(g) 1 mark for correct species and state symbols 1 mark for balance							B1 × 2	
6 (c)	6NaOH(aq) + Fe₂(SO₄)₃(aq) → 2Fe(OH)₂(s) + 3Na₂SO₄(aq) 1 mark for correct species and state symbols 1 mark for balance							B1 × 2	

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