

Experiment 30

Advance Study Assignment: Determination of Iron by Reaction with Permanganate — A Redox Titration

1. Write the balanced net ionic equation for the reaction between MnO_4^- ion and Fe^{2+} ion in acidic solution.



2. How many moles of Fe^{2+} ion can be oxidized by 1.85×10^{-3} moles MnO_4^- ion in the reaction in Question 1?

$$5\text{Fe}^{2+} \times \frac{1.85 \times 10^{-3} \text{ moles MnO}_4^-}{1}$$

$$\underline{9.25 \times 10^{-3} \text{ moles}}$$

3. A solid sample containing some Fe^{2+} ion had a total mass of 0.9791 g. It required 18.2 mL of 0.02304 M KMnO_4 to titrate the Fe^{2+} in the dissolved sample to a pink endpoint.

- a. How many moles of MnO_4^- ion were required?

$$0.02304 \text{ M} \times 18.2 = 0.419328$$

$$\underline{0.4193 \times 10^{-3} \text{ moles}}$$

- b. How many moles of Fe^{2+} were present in the sample?

$$5 \text{ moles of Fe}^{2+} \times 0.4193 \times 10^{-3} = 2.0965$$

$$\underline{2.0965 \times 10^{-3} \text{ moles}}$$

(continued on following page)

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In a global defined by information and interconnectivity, the enchanting power of words has acquired unparalleled significance. Their ability to kindle emotions, provoke contemplation, and ignite transformative change is truly awe-inspiring. Enter the realm of "**Advance Study Assignment Determination Of Iron Answers**," a mesmerizing literary masterpiece penned by way of a distinguished author, guiding readers on a profound journey to unravel the secrets and potential hidden within every word. In this critique, we shall delve in to the book is central themes, examine its distinctive writing style, and assess its profound impact on the souls of its readers.

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