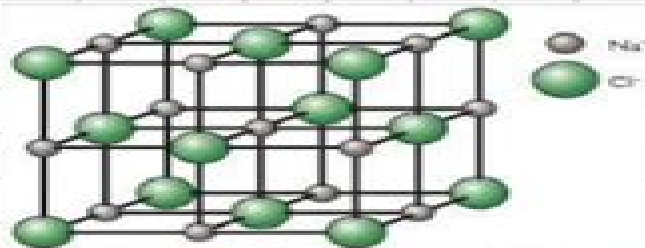


Ionic Bonding

(Between Metals and Non Metals : Group 1,2,6,7)

Ionic Bond: Bond formed by the electrostatic attraction between oppositely charged ions



Giant Lattice Structure
(exploded view)

EVIDENCE FOR THE EXISTENCE OF IONS

- 1) Electrolysis of aqueous ionic compound or a solution of it
- 2) Potassium Permanganate experiment

A) Experiment 1 : Electrolysis

- (a) • CuCrO_4 is dissolved in HCl acid. It is saturated with urea which increases the density of the solution.
- It is poured into a U tube, the liquid is originally green but as the electricity is applied the color shifts up the tube as the ions migrate towards the electrodes.

* $\text{CuCrO}_4 \rightarrow \text{Green}$

* -ve electrode $\rightarrow$ Blue Cu^{2+} ions

* +ve electrode $\rightarrow$ Yellow CrO_4^{2-} ions

- (b) • NaCl in aqueous form is poured into the U-tube and electricity is passed through it.

* Aqueous $\text{NaCl} \rightarrow$ Colorless

* -ve electrode $\Rightarrow$ Solid Na forms

* +ve electrode $\rightarrow$ Cl_2 gas is given off that can bleach damp litmus paper

Advanced Ionic Bonding Chem Quest 20

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