# Calculations in Analytical Chemistry

### The Mole

The mole is the SI unit for amount of substance. It corresponds to the amount of a chemical substance that contains the same number of particles present in 12 grams of C-12 which is equivalent to Avogadro's number, that is, NA=6.022X10^23.

The molar mass is given as the mass of a compound in grams per mole of that substance. It is calculated by adding the molar masses of all the component elements in the compound. Take glucose for example:

Element	Molar mass per element	Number of atoms	Mass
С	12 g/mol	6	72 g/mol
н	1 g/mol	12	12g/m
0	16 g/mol	6	96 g/mol
Molar Mass			180 g/mol

### Mole and Millimole Calculations

The millimole is simply 1/1000 of a mole. It is sometimes used in calculations because of its convenience.

Example (Skoog, et al., 2014):

 Find the number of moles and millimoles of benzoic acid (122.1 g/mol) that are contained in 2.00 g of pure acid.

If we use HBz to represent benzoic acid:

$$HBz_{mol} = 2.00 \frac{g_{HBz}}{g_{HBz}} \times \frac{1mol_{HBz}}{122.1 \frac{g_{HBz}}{g_{HBz}}} = 0.0164 mol_{HBz}$$

To obtain the number of milimoles, the millimolar mass of benzoic acid will be used:

$$HBz_{mol} = 2.00 \frac{g_{HBz}}{g_{HBz}} \times \frac{1mmol_{HBz}}{0.1221 \frac{g_{HBz}}{g_{HBz}}} = 16.4mmol_{HBz}$$

 What is the mass in grams of Na\* (22.99 g/mol) in 25.0 g of Na<sub>2</sub>SO<sub>4</sub> (142.0 g/mol)?

$$mass_{Na^{+}} = 25.0 \frac{g_{Na_{1}50_{4}}}{142.0 \frac{g_{Na_{1}50_{4}}}{142.0 \frac{g_{Na_{1}50_{4}}}{1400 \frac{g_{Na_{1}50_{4}}}{1000 \frac{g_{Na_{1}50_{4}}}}} \times \frac{2 \frac{mol_{Na^{+}}}{1000 \frac{g_{Na_{1}50_{4}}}}}{1 \frac{mol_{Na^{+}}}{1000 \frac{g_{Na_{1}50_{4}}}}} \times \frac{22.99 \frac{g_{Na^{+}}}{1000 \frac{g_{Na_{1}50_{4}}}}}{1 \frac{mol_{Na^{+}}}{1000 \frac{g_{Na_{1}50_{4}}}}} \times \frac{22.99 \frac{g_{Na^{+}}}{1000 \frac{g_{Na_{1}50_{4}}}}}{1 \frac{g_{Na^{+}}}{1000 \frac{g_{Na_{1}50_{4}}}}} \times \frac{20.99 \frac{g_{Na^{+}}}{1000 \frac{g_{Na_{1}50_{4}}}}}{1 \frac{g_{Na_{1}50_{4}}}{1000 \frac{g_{Na_{1}50_{4}}}}} \times \frac{20.99 \frac{g_{Na^{+}}}{1000 \frac{g_{Na_{1}50_{4}}}{1000 \frac{g_{Na_{1}$$

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## **Table of Contents Analytical Chemistry Calculation Questions And Answers**

- 1. Understanding the eBook Analytical Chemistry Calculation Questions And Answers
  - The Rise of Digital Reading Analytical Chemistry Calculation Questions And Answers
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Analytical Chemistry Calculation Questions And Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Analytical Chemistry Calculation Questions And Answers
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Analytical Chemistry Calculation Questions And Answers
  - Personalized Recommendations
  - Analytical Chemistry Calculation Questions And Answers User Reviews and Ratings
  - Analytical Chemistry Calculation Questions And Answers and Bestseller Lists
- 5. Accessing Analytical Chemistry Calculation Questions And Answers Free and Paid eBooks

- Analytical Chemistry Calculation Questions And Answers Public Domain eBooks
- Analytical Chemistry Calculation Questions And Answers eBook Subscription Services
- Analytical Chemistry Calculation Questions And Answers Budget-Friendly Options
- 6. Navigating Analytical Chemistry Calculation Questions And Answers eBook Formats
  - o ePub, PDF, MOBI, and More
  - o Analytical Chemistry Calculation Questions And Answers Compatibility with Devices
  - Analytical Chemistry Calculation Questions And Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Analytical Chemistry Calculation Questions And Answers
  - Highlighting and Note-Taking Analytical Chemistry Calculation Questions And Answers
  - Interactive Elements Analytical Chemistry Calculation Questions And Answers
- 8. Staying Engaged with Analytical Chemistry Calculation Questions And Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Analytical Chemistry Calculation Questions And Answers
- 9. Balancing eBooks and Physical Books Analytical Chemistry Calculation Questions And Answers
  - Benefits of a Digital Library
  - o Creating a Diverse Reading Collection Analytical Chemistry Calculation Questions And Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Analytical Chemistry Calculation Questions And Answers
  - Setting Reading Goals Analytical Chemistry Calculation Questions And Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Analytical Chemistry Calculation Questions And Answers
  - Fact-Checking eBook Content of Analytical Chemistry Calculation Questions And Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development

- Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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