MIXED GAS LAWS WORKSHEET - SOLUTIONS

 How many moles of gas occupy 98 L at a pressure of 2.8 atmospheres and a temperature of 292 K?

n = <u>PV</u> = <u>(2.8 atm)(98 L)</u> = 11 moles of gas (0.0821 Latm/mol.K)(292 K)

2) If 5.0 moles of O₂ and 3.0 moles of N₂ are placed in a 30.0 L tank at a temperature of 25° C, what will the pressure of the resulting mixture of gases be? 25° C = 298 K

- $O_{22} P = \frac{nRT}{V} = \frac{(5.0 \text{ mol})(0.0821 \text{ Latm/mol.K})(298 \text{ K})}{(30.0 \text{ L})} = 4.1 \text{ atm}$
- N₂: $P = \frac{nRT}{V} = \frac{(3.0 \text{ mol})(0.0821 \text{ Latm/molK})(298 \text{ K})}{(30.0 \text{ L})} = 2.4 \text{ atm}$

 $P_{TM} = P_{02} + P_{N2} = 4.1$ atm + 2.4 atm = 6.5 atm Alternately you can substitute total moles of gas (8.0) as the total pressure depends on the

total moles of gas – what type of gas is irrelevant

3) A balloon is filled with 35.0 L of helium in the morning when the temperature is 20.0°C. By noon the temperature has risen to 45.0°C. What is the new volume of the balloon?

 $T_1 = 20.0^{\circ} C = 293 \text{ K}, V_2 = 35.0 \text{ L}, T_2 = 45.0^{\circ} C = 318 \text{ K}, V_2 = 7$ $V_2 = \frac{V_1 T_2}{T_1} = \frac{(35.0 \text{ L})(318 \text{ K})}{(293 \text{ K})} = 38.0 \text{ L}$

4) A 35 L tank of oxygen is at 315 K with an internal pressure of 190 atmospheres. How many moles of gas does the tank contain?

n = <u>PV</u> = <u>(190 atm)(35 L)</u> = 260 moles of gas RT (0.0821 Latm/mol.K)(315 K)

5) A balloon that can hold 85 L of air is inflated with 3.5 moles of gas at a pressure of 1.0 atmosphere. What is the temperature in °C of the balloon?

T = PV = (1 atm)(85 L) = 296 K = 23°C nR (3.5 mol)(0.0821 Latm/mol.K)

6) CaCO₃ decomposes at 1200°C to form CO₂ gas and CaO. If 25 L of CO₂ are collected at 1200°C, what will the volume of this gas be after it cools to 25°C?

 $T_1 = 1200$ °C = 1473 K, $V_1 = 25$ L, $T_2 = 25$ °C = 298 K, $V_2 = 7$ $V_2 = V_1 T_2 = (25 \text{ L})(298 \text{ K}) = 5.1 \text{ L}$ (1473 K)

7) A helium balloon with an internal pressure of 1.00 atm and a volume of 4.50 L at 20.0°C is released. What volume will the balloon occupy at an altitude where the pressure is 0.600 atm and the temperature is -20.0°C?

 P_1 = 1.00 atm, V_1 = 4.50 L, T_1 = 20.0°C = 293 K, P_2 = 0.600 atm, V_2 = 7, T_2 = -20.0°C = 253 K V_2 = $\frac{P_1 V_1 T_2}{I_1 P_2}$ = $\frac{(1.00 \text{ atm})(4.50 \text{ L})(253 \text{ K})}{(293 \text{ K})(0.600 \text{ atm})}$ = 6.48 L

Answer Sheet To Mixed Gas Laws

Tobias Bleicker

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