

## Chemistry Unit 8 Worksheet 4

### Samples of Every Kind of Problem

*On a separate sheet of paper, write a complete solution to each of the problems below. Follow the procedure outlined in class. Be sure to circle your final answer.*

1. Calculate the number of moles of potassium chlorate,  $\text{KClO}_3$  (s), that must decompose to produce potassium chloride,  $\text{KCl}$  (s), and 1.8 moles of oxygen gas.

|           |                                                           |      |      |                                     |
|-----------|-----------------------------------------------------------|------|------|-------------------------------------|
| Equation: | $2 \text{KClO}_3 \rightarrow 2 \text{KCl} + 3 \text{O}_2$ |      |      | 1.2 moles<br>$\text{KClO}_3$ needed |
| Before:   | 1.2                                                       | 0    | 0    |                                     |
| Change:   | -1.2                                                      | +1.2 | +1.8 |                                     |
| After:    | 0                                                         | 1.2  | 1.8  |                                     |

2. In a single displacement reaction, magnesium metal reacts with hydrochloric acid to produce magnesium chloride and hydrogen gas. How many moles of hydrochloric acid are needed to completely react with 2.43 g of magnesium?

|           |                                                                   |          |         |       |                                    |
|-----------|-------------------------------------------------------------------|----------|---------|-------|------------------------------------|
| Equation: | $\text{Mg} + 2 \text{HCl} \rightarrow \text{MgCl}_2 + \text{H}_2$ |          |         |       | 0.200 moles<br>$\text{HCl}$ needed |
| Before:   | 0.100                                                             | ?        | 0       | 0     |                                    |
| Change:   | -0.100                                                            | -0.200 + | 0.100 + | 0.100 |                                    |
| After:    | 0                                                                 | ?        | 0.100   | 0.100 |                                    |

3. Ethane,  $\text{C}_2\text{H}_6$  reacts with oxygen gas to produce carbon dioxide gas and water vapor. What mass of oxygen gas is required to react with 2.20 moles of ethane?

|           |                                                                                          |       |       |       |
|-----------|------------------------------------------------------------------------------------------|-------|-------|-------|
| Equation: | $2 \text{C}_2\text{H}_6 + 7 \text{O}_2 \rightarrow 4 \text{CO}_2 + 6 \text{H}_2\text{O}$ |       |       |       |
| Before:   | 2.20                                                                                     | xs    | 0     | 0     |
| Change:   | -2.20                                                                                    | -7.70 | +4.40 | +6.60 |
| After:    | 0                                                                                        | xs    | 4.40  | 6.60  |

$$\text{O}_2 \text{ needed: } 7.70 \text{ mole} \times \frac{32.0 \text{ g}}{1 \text{ mole}} = 246 \text{ g O}_2 \text{ required}$$

4. Determine the mass of sodium nitrate produced when 0.73 g of nickel (II) nitrate reacts with sodium hydroxide according to the following equation:



$$0.73 \text{ g} \times \frac{1 \text{ mole}}{183 \text{ g}} = .0040 \text{ mole}$$

|           |                                                                                                 |         |         |         |
|-----------|-------------------------------------------------------------------------------------------------|---------|---------|---------|
| Equation: | $\text{Ni}(\text{NO}_3)_2 + 2 \text{NaOH} \rightarrow \text{Ni}(\text{OH})_2 + 2 \text{NaNO}_3$ |         |         |         |
| Before:   | 0.0040                                                                                          | XS      | 0       | 0       |
| Change:   | -0.0040                                                                                         | -0.0080 | +0.0040 | +0.0080 |
| After:    | 0                                                                                               | XS      | 0.0040  | 0.0080  |

$$\text{NaNO}_3 \text{ produced: } 0.0080 \text{ mole} \times \frac{85.0 \text{ g}}{1 \text{ mole}} = 0.68 \text{ g}$$

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**About the Author** DR DENNIS ENGLIN earned his bachelor s from Westmont College his master of science from California State University and his EdD from the University of Southern California He enjoys teaching animal biology vertebrate biology wildlife biology organismic biology and astronomy at The Master s University His professional memberships include the Creation Research Society the American Fisheries Association Southern California Academy of Sciences Yellowstone Association and Au Sable Institute of Environmental Studies

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