

Experiment 10: Electrochemistry with a Standard Solution from: *Basic Laboratory Experiments*, Experiment 10, 45-50

The purpose of this experiment is to determine the standard reduction potential for several half-reactions. In this experiment, a half-cell is constructed from a metal electrode immersed in a solution of its ions. The standard reduction potential for this half-cell is determined by measuring the cell potential of a cell constructed from this half-cell and a standard hydrogen electrode. Other standard half-cells, such as the standard calomel electrode and the standard silver-silver chloride electrode, are also used in this experiment. The standard reduction potential for the half-cell containing the standard hydrogen electrode is defined as 0.00 V.

In this experiment, you will determine the standard reduction potential for several half-cells and compare your results with the standard reduction potentials listed in the textbook.

Objectives:

1. to determine the standard reduction potential for several half-cells
2. to compare the standard potential of the half-cells determined in this experiment with the standard reduction potentials listed in the textbook
3. to determine the standard potential for several half-cells
4. to determine the standard potential of a half-cell containing a standard solution
5. to compare the standard potential of the half-cells determined in this experiment with the standard reduction potentials listed in the textbook

Materials:

- standard solution
- standard solution (0.1M)
- standard solution
- standard solution
- standard solution
- standard solution
- standard solution
- standard solution
- standard solution
- standard solution

Procedure:

1. Prepare standard solution
2. Prepare standard solution
3. Prepare standard solution
4. Prepare standard solution

Results:

1. Calculate the standard potential for each half-cell.
2. Compare the standard potential for each half-cell with the standard reduction potential listed in the textbook.

Discussion:

1. Discuss the results of the experiment.
2. Discuss the standard reduction potential for each half-cell and compare it with the standard reduction potential listed in the textbook.

Answers For Lab 6b Heath Chemistry

Tao Wei



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