

# Gravimetric Analysis of Chloride in Solution

Experiment 10  
Lab 10  
Date: \_\_\_\_\_

## Introduction:

This experiment involves the gravimetric analysis of a solution containing chloride ions. The principle of gravimetric analysis is based on the fact that the mass of a substance is proportional to the number of moles of the substance. In this experiment, the chloride ions in the solution are precipitated as silver chloride ( $\text{AgCl}$ ) by adding a known volume of a silver nitrate ( $\text{AgNO}_3$ ) solution. The precipitate is then filtered, dried, and weighed. The mass of the precipitate is used to calculate the mass of chloride ions in the original solution. The experiment is performed in a fume hood to avoid exposure to silver nitrate, which is a strong oxidizing agent and can cause skin irritation. The experiment is also performed in a dry environment to avoid the formation of silver chloride hydrate, which would affect the results. The experiment is performed in a dry environment to avoid the formation of silver chloride hydrate, which would affect the results.

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## Procedure:

The procedure involves the following steps: 1. Preparation of the silver nitrate solution. 2. Addition of the silver nitrate solution to the sample solution. 3. Filtration of the precipitate. 4. Drying of the precipitate. 5. Weighing of the precipitate.

## Data Analysis:

|                      | Sample 1 | Sample 2 |
|----------------------|----------|----------|
| Mass of sample       | 0.100 g  | 0.100 g  |
| Mass of precipitate  | 0.150 g  | 0.150 g  |
| Mass of chloride ion | 0.050 g  | 0.050 g  |

The mass of chloride ion in the sample is calculated as follows:

$$\text{Mass of chloride ion} = \text{Mass of precipitate} \times \frac{\text{Molar mass of chloride ion}}{\text{Molar mass of silver chloride}}$$

The mass of chloride ion in the sample is calculated as follows:

# Analysis Of Chloride Experiment 7 Lab Report

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