

# Prelab

## Enzymatic Decomposition and Analysis of Hydrogen Peroxide

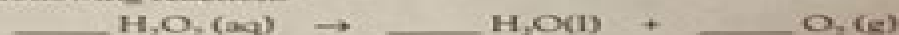
Name \_\_\_\_\_

Date \_\_\_\_\_

Section \_\_\_\_\_

1. What gas is being collected in today's experiment?

2. Balance the following reaction:



3. In today's experiment you determine that you collected  $1.33 \times 10^{-3}$  mols of oxygen. Calculate (show your setup for the conversion!) how many mols of  $\text{H}_2\text{O}_2$  were in your hydrogen peroxide solution. Note: You will need a conversion factor from the equation you balanced in question 2.

\_\_\_\_\_ mols of  $\text{H}_2\text{O}_2$

4. Based on your answer above, how many grams of  $\text{H}_2\text{O}_2$  were in the hydrogen peroxide solution? ( $\text{H}_2\text{O}_2$  has a molecular weight of 34.0 g/mol)

\_\_\_\_\_ grams of  $\text{H}_2\text{O}_2$

5. Assume the grams of  $\text{H}_2\text{O}_2$  calculated in question 4 were found in a 6.00 gram sample of a hydrogen peroxide solution. Calculate the mass percent  $\text{H}_2\text{O}_2$  in the solution.

\_\_\_\_\_ %  $\text{H}_2\text{O}_2$

# Analysis Of Hydrogen Peroxide Prelab Answers

**Ralph Petrucci, William  
Harwood, Geoffrey Herring**



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