AP CHEMISTRY MULTIPLE CHOICE OUESTIONS Unit 8 - Introduction to Equilibrium

 $N_2O_4(g) \rightleftharpoons 2NO_2(g) + \frac{1}{2}O_2(g)$ 1.0

> The equilibrium constant for the gas phase reaction above is 95 at 25°C. What is the value of the ourilibrium constant for the following reaction BE 25 YEAR

- CAU (95)2
- GE1 95

3... $C_*H_*(g) + H_*(g) \xrightarrow{\text{plainers}} C_*H_*(g)$

C₂H₂(g) is reduced by H₂(g) in the presence of a solid platinum catalyst, as represented by the equation above. Factors that could affect the rate of the reaction include which of the following?

- Changes in the partial pressure of H₂(g).
- II. Changes in the particle size of the platinum.
- III. Changes in the temperature of the reaction. SALES SECTION
- (A) III only
- (B) I and II only
- (C) I and III only
- (D) III and III only (Ellip L. III., amed IIII.

shut+ 2 Bao(4) = 2 Bao(4) + O(4) AH = 162 kumol en dotharmic

A sealed rigid vessel contains BaO₂(x) in equilibrium with BaO(s) and O₂(g) as represented by the equation above. Which of the following changes will increase the amount of BaO₇(s) in the vessel?

- (A) Removing a small amount of O₂(x) | K
- (B) Removing a small amount of BaO(x) X
- (C) Adding He gas to the vessel X
- (D) Lowering the temperature \

- For which of the following reaction mixtures at equilibrium would reducing the volume of the container at constant temperature cause the equilibrium to shift toward the products? [MANALAR (A) SrCO₂(t) = SrO(1) + CO₂(g)
 - (B) C(c) + CO(c) = 2 CO(c)
 - (C) CO(g) + H₂O(g) = CO₂(g) + H₂(g)
 - DEROUS = POLED + CILED DEROUS = NOVED + CONSTRUCT GOLD

X(g) + Y(g) = 2Z(g)

When 4.00 mot each of X(g) and Y(g) are placed in a 1.00 L vessel and allowed to react at constant temperature according to the equation above, 6.00 mol of Z(g) is produced. What is the value of the equilibrium constant, K_c ?

CA3: 3 CERT 16-

The synthesis of $CH_2OH(g)$ from CO(g) and $H_2(g)$ is represented by the equation above. The value of K

> for the reaction at 483 K is 14.5. increase heat shifts le Which of the following explains the effect on the equilibrium constant. K,., when the temperature of the reaction system is increased to 650 K?

- (A) K, will increase because the activation. energy of the forward reaction increases. more than that of the reverse reaction.
- (B) K, will increase because there are more
- reactant molecules than product molecules. (C) K, will decrease because the reaction is exothermic.
- (D) K, is constant and will not change.

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Lei Shi

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AP Chemistry Premium, 2024: 6 Practice Tests + Comprehensive Review + Online Practice Neil D. Jespersen, Pamela Kerrigan, 2023-07-04 A guide to taking the Advanced Placement exam in chemistry featuring a review of major chemistry concepts practice and diagnostic tests test taking strategies an overview of the test and practice problems AP Chemistry Premium, 2022-2023: Comprehensive Review with 6 Practice Tests + an Online Index Medicus ,1997 Timed Test Option Neil D. Jespersen, Pamela Kerrigan, 2021-07-06 Be prepared for exam day with Barron's Trusted content from AP experts Barron's AP Chemistry Premium 2022 2023 includes in depth content review and online practice It's the only book you ll need to be prepared for exam day Written by Experienced Educators Learn from Barron s all content is written and reviewed by AP experts Build your understanding with comprehensive review tailored to the most recent exam Get a leg up with tips strategies and study advice for exam day it s like having a trusted tutor by your side Be Confident on Exam Day Sharpen your test taking skills with 6 full length practice tests 3 in the book and 3 more online Strengthen your knowledge with in depth review covering all Units on the AP Chemistry Exam Reinforce your learning with practice questions at the end of each chapter Interactive Online Practice Continue your practice with 3 full length practice tests on Barron s Online Learning Hub Simulate the exam experience with a timed test option Deepen your understanding with detailed answer explanations and expert advice Gain confidence with automated scoring to check your learning progress **Subject** Environmental Health Perspectives ,2003-07 Guide to Books in Print ,2001 **Applied Science & Technology** Methods in Ecosystem Science Osvaldo E. Sala, Robert B. Jackson, Harold A. Mooney, Robert W. **Index** .1996 Howarth, 2013-12-01 Ecology at the ecosystem level has both necessitated and benefited from new methods and technologies as well as those adapted from other disciplines With the ascendancy of ecosystem science and management the need has arisen for a comprehensive treatment of techniques used in this rapidly growing field Methods in Ecosystem Science answers that need by synthesizing the advantages disadvantages and tradeoffs associated with the most commonly used techniques in both aquatic and terrestrial research The book is divided into sections addressing carbon and energy dynamics nutrient and water dynamics manipulative ecosystem experiements and tools to synthesize our understanding of ecosystems Detailed information about various methods will help researchers choose the most appropriate methods for their particular studies Prominent scientists discuss how tools from a variety of disciplines can be used in ecosystem science at different scales Current Index to Journals in Education ,1998 2024-25 NCERT Class-XI & XII Chemistry Solved Papers YCT Expert Team, 2024 25 NCERT Class XI XII Chemistry Solved Papers 608 1195 E This book contains previous solved papers and 6070 solved objective questions with detail explanation **Scientific Discovery Processes in Humans and Computers** Morton Wagman, 2000-05-30 Wagman offers a critical analysis of current theory and research in the psychological and

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