

ANION ANALYSIS

OBJECTIVES FOR THE EXPERIMENT

The student will be able to do the following:

1. Write a general description of the qualitative analysis technique for identification of unknown substances.
2. Plan a procedure for the analysis of an unknown solution of an anion and describe that procedure with a flow chart.
3. Identify different anions using a qualitative analysis technique.

BACKGROUND

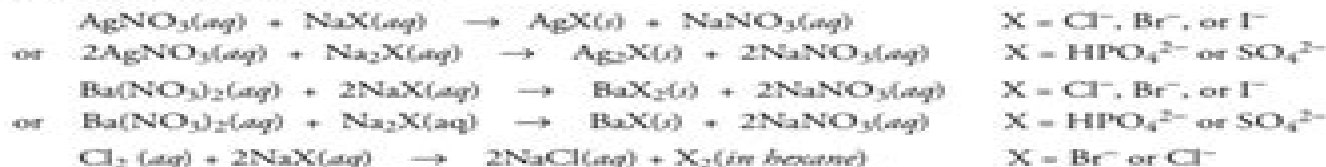
There are two main branches of analytical chemistry, qualitative analysis and quantitative analysis. Qualitative analysis involves the experimental determination of the nature of the substances in a sample. It answers the question, what's in there? Quantitative analysis involves the determination of how much of a substance is present in a sample. A chemist might test the water downstream from an out-fall of a chemical plant. The first step would be to do qualitative analysis techniques to determine whether it contains pollutants. The next step would be to use quantitative analysis techniques to find out if the concentrations were high enough to cause concern.

The goal of this experiment is to develop techniques for identifying unknown substances. You will identify unknown solutions as containing one of the following: Cl^- , Br^- , I^- , HPO_4^{2-} , or SO_4^{2-} .

You will do the following general steps.

1. You will first do tests with solutions that you know contain one of the anions.
 - a. You will determine how the anions react with silver nitrate and barium nitrate.
 - b. You will determine whether precipitates dissolve in acidic solutions.
 - c. You will determine how Br^- and I^- react with chlorine, Cl_2 .
2. Based on the observations made on these "knowns", you can devise a plan to analyze each of your unknowns. You will describe this plan with a flow chart. See below.
3. You will then do tests on three unknown solutions to determine which anion is in each solution.

The general reactions in this experiment consist of two double displacement reactions and one oxidation-reduction reaction as shown below:

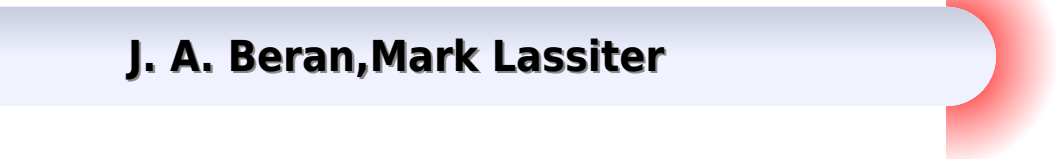


In the silver nitrate and barium nitrate tests, you will watch for the possible formation of the precipitates AgX , Ag_2X , BaX_2 , or BaX . The colors of these precipitates can be quite helpful. The appearance of any solids in these tests will show that a reaction has occurred. Even a slight cloudiness is an indication of a positive test.

If a precipitate forms in a silver nitrate or barium nitrate test, you will add nitric acid to see if the ionic compound will dissolve. If the anion is basic, the ionic compound should dissolve. When a solid ionic compound precipitates from a solution, the ions do not remain permanently in the ionic solid. Ions are constantly escaping from the solid into solution, and ions are constantly colliding with the solid and returning to the solid. These two changes occur at the same rate, so there is not normally any net change in the amount of solid in

Anion Analysis Lab Report

J. A. Beran, Mark Lassiter



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Table of Contents Anion Analysis Lab Report

1. Understanding the eBook Anion Analysis Lab Report
 - The Rise of Digital Reading Anion Analysis Lab Report
 - Advantages of eBooks Over Traditional Books
2. Identifying Anion Analysis Lab Report
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Anion Analysis Lab Report
 - User-Friendly Interface
4. Exploring eBook Recommendations from Anion Analysis Lab Report
 - Personalized Recommendations
 - Anion Analysis Lab Report User Reviews and Ratings
 - Anion Analysis Lab Report and Bestseller Lists

5. Accessing Anion Analysis Lab Report Free and Paid eBooks
 - Anion Analysis Lab Report Public Domain eBooks
 - Anion Analysis Lab Report eBook Subscription Services
 - Anion Analysis Lab Report Budget-Friendly Options
6. Navigating Anion Analysis Lab Report eBook Formats
 - ePub, PDF, MOBI, and More
 - Anion Analysis Lab Report Compatibility with Devices
 - Anion Analysis Lab Report Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Anion Analysis Lab Report
 - Highlighting and Note-Taking Anion Analysis Lab Report
 - Interactive Elements Anion Analysis Lab Report
8. Staying Engaged with Anion Analysis Lab Report
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Anion Analysis Lab Report
9. Balancing eBooks and Physical Books Anion Analysis Lab Report
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Anion Analysis Lab Report
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Anion Analysis Lab Report
 - Setting Reading Goals Anion Analysis Lab Report
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Anion Analysis Lab Report
 - Fact-Checking eBook Content of Anion Analysis Lab Report
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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