

# Pre-lab Questions for Experiment #5

## Rates of Chemical Reactions I: A Clock Reaction

1) Complete and balance the following reactions:



2) The rate law for the reaction



is  $\text{rate} = k[\text{H}_2][\text{NO}]^2$ . Which of the following mechanisms can be ruled out on the basis of the observed rate expression?

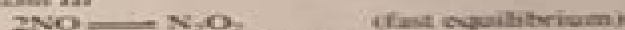
*Mechanism I*



*Mechanism II*



*Mechanism III*



3) Consider the reaction



From the following data obtained at a certain temperature, determine the order of the reaction and calculate the rate constant:

| [A] (M) | [B] (M) | Rate (M/s)            |
|---------|---------|-----------------------|
| 1.50    | 1.50    | $3.20 \times 10^{-1}$ |
| 1.50    | 2.50    | $3.20 \times 10^{-1}$ |
| 3.00    | 1.50    | $6.40 \times 10^{-1}$ |

4) The following gas-phase reaction was studied at  $290^\circ\text{C}$  by observing the change in pressure as a function of time in a constant-volume vessel:



Determine the order of the reaction and the rate constant based on the following data:

| Time (s) | Partial Pressure of $\text{ClCO}_2\text{CCl}_3$ (mmHg) |
|----------|--------------------------------------------------------|
| 0        | 15.76                                                  |
| 181      | 12.64                                                  |
| 513      | 8.73                                                   |
| 1164     | 4.44                                                   |

# Answer Of Pre Lab Rates Of Reaction

**Rosemary A. Marusak, Kate Doan, Scott  
D. Cummings**

## Answer Of Pre Lab Rates Of Reaction:

*Integrated Approach to Coordination Chemistry* Rosemary A. Marusak, Kate Doan, Scott D. Cummings, 2007-03-07  
Coordination chemistry is the study of compounds formed between metal ions and other neutral or negatively charged molecules. This book offers a series of investigative inorganic laboratories approached through systematic coordination chemistry. It not only highlights the key fundamental components of the coordination chemistry field, it also exemplifies the historical development of concepts in the field. In order to graduate as a chemistry major that fills the requirements of the American Chemical Society, a student needs to take a laboratory course in inorganic chemistry. Most professors who teach inorganic chemistry laboratory prefer to emphasize coordination chemistry rather than attempting to cover all aspects of inorganic chemistry because it keeps the students focused on a cohesive part of inorganic chemistry which has applications in medicine, the environment, molecular biology, organic synthesis, and inorganic materials. **Lab Manual Chemistry Class XII -by Dr. K. N. Sharma, Dr. Subhash Chandra Rastogi, Er. Meera Goyal (SBPD Publications)** Dr. K. N. Sharma, Dr. Subhash Chandra Rastogi, Er. Meera Goyal, 2021-07-03. Highly Useful for Various Engineering and Medical Competitive Examinations. Chemistry Lemay, Prentice-Hall Staff, 1996. **Nuclear Science Abstracts**, 1971-05. Chemical Investigations Nancy Konigsberg Kerner, 1986. **The Latest and Best of TESS**, 1991. **Chemistry** McGraw-Hill Staff, 2001-07. **Experimental Organic Chemistry** Daniel R. Palleros, 2000-02-04. This cutting edge lab manual takes a multiscale approach presenting both micro, semi micro, and macroscale techniques. The manual is easy to navigate with all relevant techniques found as they are needed. Cutting edge subjects such as HPLC, bioorganic chemistry, multistep synthesis, and more are presented in a clear and engaging fashion. **Who's the New Kid in Chemistry?** John D. Butler, 2013-12-12. Who's the New Kid in Chemistry offers an unprecedented look at student engagement and teacher best practices through the eyes of an educational researcher enrolled as a public high school student. Over the course of seventy-nine consecutive days, John D. Butler participates in and observes Rhode Island 2013 Teacher of the Year Jessica M. Waters's high school chemistry class, documenting his experiences as they unfold. Who's the New Kid in Chemistry is a compelling example of what can be accomplished when an educational researcher and teacher collaborate in the classroom. This work includes a discussion on flexible homework assignments, data-driven instruction, and thirty teacher best practices. This book is an invaluable resource for teachers across all content areas, masters and doctoral research method classes, and future Teachers of the Year. **Laboratory Manual on Biotechnology** P. M. Swamy, 2008. Selected Water Resources Abstracts, 1988-04. **Chemical Engineering Education**, 2002. **Pre-lab Exercises for Experimental Organic Chemistry** Royston M. Roberts, John C. Gilbert, Stephen F. Martin, 1994. General Chemistry Hunt, 1976-06-07. *Laboratory Experiments for General Chemistry* Harold R. Hunt, Toby F. Block, George M. McKelvy, 1998. Methods in Biotechnology Seung-Beom Hong, M. Bazlur Rashid, Lory Z. Santiago-Vázquez, 2016-08-01. As rapid advances in biotechnology occur, there is a need for a

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