Hydrolysis -The Reaction of lons with Water

Lab Report

Purpose

The purpose of this lab is to predict whether a substance is acidic, basic, or neutral and to use the universal indicator solution to test it.

Hypothesis

Sodium Acetate	$NaCH_3COO_{(a)} \rightarrow Na^*_{(aq)} + CH_3COO^*_{(aq)}$ $CH_3COO^*_{(aq)} + H(OH)_{(b} \leftrightarrow CH_3COOH_{(aq)} + OH^*_{(aq)}$		Basic
Sodium Chloride	NaCl _{so} ++Na* _{ted} + Cf _{ted}		Neutral
Ammonium Chloride	$NH_4Cl_{03}\rightarrow NH_{4(eq)}+Cl_{(eq)}$ $NH_4^*_{(eq)}+H(OH)_{(3}\leftrightarrow NH_4(OH)_{(eq)}+H^*_{(eq)}$		Acidic
Ammonium Sulfate	$(NH_4)_2SO_{4(g)} \rightarrow NH_4^*_{(acg)} + SO_4^{2^*}_{(acg)}$ $NH_4^*_{(acg)} + H(OH)_{(g)} \leftrightarrow NH_4(OH)_{(acg)} + H^*_{(acg)}$ $SO_4^{2^*}_{(acg)} + H(OH)_{(g)} \leftrightarrow HSO_4^*_{(acg)} + OH^*_{(acg)}$	Ka=5,4×10 ⁻¹⁰	Acidic
		Kb=8.3×10 ⁻¹³	
Aluminum Chloride	$AICI_{(eq)} \rightarrow AIS^*_{(eq)} + CI^*_{(eq)}$ $AIS^*_{(eq)} + H(OH)_{(eq)} \rightarrow AI(OH)_{(eq)} + H^*_{(eq)}$		Acidic
Calcium Nitrate	Ca(NO ₃) _{Nel} →Ca ^{2*} _(leg) + No _{Neg}		Neutral
Iron(III) Sulfate	$\begin{array}{l} {\sf Fe}_2({\sf SO}_4)_{3(q)} {\to} {\sf Fe}^{3r}_{(aq)} + {\sf SO}_4^{2r}_{(aq)} \\ {\sf Fe}^{3r}_{(aq)} + {\sf H}({\sf OH})_{(3} {\leftrightarrow} {\sf Fe}({\sf OH})^{2r}_{(aq)} + {\sf H}^*_{(aq)} \\ {\sf Fe}^{3r}_{(aq)} + {\sf H}({\sf OH})_{(0} {\leftrightarrow} {\sf HFe}^{4r}_{(aq)} + {\sf OH}^*_{(aq)} \\ {\sf SO}_4^{2r}_{(aq)} + {\sf H}({\sf OH})_{(3} {\leftrightarrow} {\sf HSO}_4^{r}_{(aq)} + {\sf OH}^*_{(aq)} \end{array}$	Ka=6.0×10 ⁻⁸	Acidic
		Kb=1.7×10 ⁻¹²	
		Kb=8.3×10 ⁻¹³	
Sodium Carbonate	$Na_2CO_{3(q)} \rightarrow Na^*_{(q)} + CO_{3(q)}$ $CO_{3(q)} + H(OH)_{13} \leftrightarrow HCO_{3(q)} + OH_{(q)}$		Basic

Answer Of Heath Chemistry Lab 20

L Darling-Hammond

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