Question 1

- 1(a)(iii) Difference: With powdered CaCO₃ the graph has a steeper slope at the start of the reaction than with marble chips because the rate of reaction is faster due to increased surface area of the solid. Similarity: The total volume of CO2 produced is the same in both cases since the same mass (and therefore number of moles) of CaCO₃ was used in both experiments.
- $1(a)(iv) CaCO_{3(s)} + 2HNO_{3(aq)} \rightarrow Ca(NO_3)_{2(aq)} + H_2O_{(i)} + CO_{2(g)}$
- 1(a)(v) Moles of CaCO₃ used = (1.8/100) mol. = 0.018 mol.
 From the balanced equation for the reaction, 1 mole CaCO₃ produces 1 mol CO₂.
 Hence volume of gas produced at R.T.P. = (0.018 × 24000) cm³ = 432 cm³.
- 1(a)(vi) There were impurities in the CaCO, used.
- 1(a)(vii) Concentration of the acid used and temperature at which the reaction was carried out.
- 1(b)(i) Fe3+ present.
- 1(b)(ii) Magnesium ribbon decreases in size/dissolves. Solution becomes pale green.
- 1(b)(iii) $Fe^{2+}_{(ac)} + 2OH^{-}_{(ac)} \rightarrow Fe(OH)_{2(c)}$
- 1(b)(iv) SO₄2-(ac) SO₃2-(ac) and CO₃2-(ac) absent.
- 1(b)(v) White precipitate, which is soluble in aqueous ammonia, is formed.

Answers For January 2014 Chemistry

Kilic Bugra Kanat

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Sulthan Bathery Kerala in 2002 He Stood First in his College for B sc and M sc Chemistry He received his Ph D in Inorganic Chemistry from IIT Madras in the year 2010 His research interest includes phosphorus based ligands in synthetic inorganic chemistry and organometallic chemistry He has Published four research papers in reputed national and international journals He has more than four years of teaching experience in various engineering colleges

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