

## AP Chemistry Chapter 4 Practice Exam

D

1. The Claus reactions, shown below, are used to generate elemental sulfur from hydrogen sulfide.

How much sulfur (in grams) is produced from 21.5 grams of  $\text{O}_2$ ?

A) 97.0 g

B) 14.4 g

C) 21.5 g

D) 43.1 g

E) none of these

$$\frac{21.5 \text{ g O}_2}{32.00 \text{ g O}_2} \times \frac{1 \text{ mol O}_2}{1 \text{ mol O}_2} \times \frac{2 \text{ SO}_2}{3 \text{ O}_2} \times \frac{3 \text{ S}}{\text{SO}_2} \times \frac{32.06 \text{ g S}}{1 \text{ mol S}} = 43.1 \text{ g S}$$

B

2. Iron is produced from its ore by the reactions:

How many moles of  $\text{O}_2(g)$  are needed to produce 4.6 moles of  $\text{Fe}(s)$ ?A) 2.3 mole  $\text{O}_2$ B) 3.5 mole  $\text{O}_2$ C) 4.6 mole  $\text{O}_2$ D) 6.9 mole  $\text{O}_2$ E) 13.8 mole  $\text{O}_2$ 

$$\frac{4.6 \text{ mol Fe}(s)}{2 \text{ Fe}} \times \frac{3 \text{ CO}}{2 \text{ Fe}} \times \frac{1 \text{ O}_2}{2 \text{ CO}} = 3.45 \text{ mol O}_2$$

B

3. Suppose the reaction
- $\text{Ca}_3(\text{PO}_4)_2 + 3\text{H}_2\text{SO}_4 \rightarrow 3\text{CaSO}_4 + 2\text{H}_3\text{PO}_4$
- is carried out starting with 153 g of
- $\text{Ca}_3(\text{PO}_4)_2$

and 76.8 g of  $\text{H}_2\text{SO}_4$ . How much phosphoric acid will be produced?

A) 76.7 g

B) 31.1 g

C) 229.8 g

D) 115.1 g

E) 96.7 g

$$\frac{153 \text{ g Ca}_3(\text{PO}_4)_2}{310.18 \text{ g Ca}_3(\text{PO}_4)_2} \times \frac{1 \text{ mol Ca}_3(\text{PO}_4)_2}{1 \text{ mol Ca}_3(\text{PO}_4)_2} \times \frac{2 \text{ H}_3\text{PO}_4}{1 \text{ Ca}_3(\text{PO}_4)_2} = 0.97 \text{ mol H}_3\text{PO}_4$$

$$\frac{76.8 \text{ g H}_2\text{SO}_4}{98.08 \text{ g H}_2\text{SO}_4} \times \frac{1 \text{ mol H}_2\text{SO}_4}{1 \text{ mol H}_2\text{SO}_4} \times \frac{2 \text{ H}_3\text{PO}_4}{3 \text{ H}_2\text{SO}_4} = 0.51 \text{ mol H}_3\text{PO}_4$$

B

- 4.
- $\text{SO}_2$
- reacts with
- $\text{H}_2\text{S}$
- as follows:

When 7.50 g of  $\text{H}_2\text{S}$  reacts with 12.75 g of  $\text{SO}_2$ , which statement applies?

A) 6.38 g of sulfur are formed.

B) 10.6 g of sulfur are formed.C) 0.0216 moles of  $\text{H}_2\text{S}$  remain.D) 1.13 g of  $\text{H}_2\text{S}$  remain.E)  $\text{SO}_2$  is the limiting reagent.

$$\frac{7.50 \text{ g H}_2\text{S}}{34.08 \text{ g H}_2\text{S}} \times \frac{1 \text{ mol H}_2\text{S}}{1 \text{ mol H}_2\text{S}} \times \frac{3 \text{ S}}{2 \text{ H}_2\text{S}} \times \frac{32.06 \text{ g S}}{1 \text{ mol S}} = 10.89 \text{ g S}$$

$$\frac{12.75 \text{ g SO}_2}{64.06 \text{ g SO}_2} \times \frac{1 \text{ mol SO}_2}{1 \text{ mol SO}_2} \times \frac{3 \text{ S}}{1 \text{ SO}_2} \times \frac{32.06 \text{ g S}}{1 \text{ mol S}} = 19.13 \text{ g S}$$

# Answers For Ap Chemistry Test Open Ended

**John K. Gilbert, Miriam Reiner, Mary  
Nakhleh**



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**Chemistry** Steven S. Zumdahl, 1993

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**Journal of Research and Development in Education**, 1986

**Applied Science & Technology Index**, 2000

*Visualization: Theory and Practice in Science Education* John K. Gilbert, Miriam Reiner, Mary Nakhleh, 2007-12-05 External representations pictures diagrams graphs concrete models have always been valuable tools for the science teacher This book brings together the insights of practicing scientists science education researchers computer specialists and cognitive scientists to produce a coherent overview It links presentations about cognitive theory its implications for science curriculum design and for learning and teaching in classrooms and laboratories

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*Chemical Warfare Bulletin*, 1943

*Journal of Applied Chemistry*, 1869

*American Druggists' Circular and Chemical Gazette*, 1892

*Resources in Education*, 1994

**Green Organic Chemistry and its Interdisciplinary Applications** Vera M. Kolb, 2017-04-21 Green Organic Chemistry and Its Interdisciplinary Applications covers key developments in green chemistry and demonstrates to students that the developments were most often the result of innovative thinking Using a set of selected experiments all of which have been performed in the laboratory with undergraduate students it demonstrates how to optimize and develop green experiments The book dedicates each chapter to individual applications such as Engineering The chemical industry The pharmaceutical industry Analytical chemistry Environmental chemistry Each chapter also poses questions at the end with the answers included By focusing on both the interdisciplinary applications of green chemistry and the innovative thinking that has produced new developments in the field this book manages to present two key messages in a manner where they reinforce each other It provides a single and concise reference for chemists instructors and students for learning about green organic chemistry and its great and ever expanding number of applications

**The Education Index**, 1990

*A Complete Crash Course in AIEEE 2011* K.K. Arora, Dinesh Khattar, Ravi Raj Dudeja,

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