

THE CHEMISTRY OF LIFE

Living matter is composed of many different elements that are made up of small units called atoms. An atom of any element has a nucleus that is surrounded by particles called electrons. Each electron has a negative charge. A nucleus usually has two types of particles, protons and neutrons. Each proton has a positive charge. Neutrons have no charge. The number of protons in a nucleus of atom is called its atomic number. When you add up the number of protons and neutrons together, you get what is called the atomic mass. The chart below shows six elements that make up most of your body.

Use this information about atoms and the facts in the chart to answer the questions that follow the chart.

Element	Element Symbol	Atomic number	Atomic mass	Percent in the human body
Hydrogen	H	1	1	9.5
Carbon	C	6	12	18.5
Nitrogen	N	7	14	3.0
Oxygen	O	8	16	65.0
Phosphorus	P	15	31	1
Sulfur	S	16	32	0.3

1. What element is the most abundant in the human body? What is its symbol?

2. What is the atomic number of the element in question 1?

3. What makes up to 18.5 percent of the body? What is its symbol?

4. Write the atomic number of phosphorous. How many protons does it have?

5. Finish the word equation

6. Atomic mass = the number of _____ + the number of _____
7. What is the atomic mass of nitrogen?

8. How many neutrons are there in the nucleus of a nitrogen atom?

9. Write the symbol of the element found in only 0.3 percent of human body.

10. Which element in the chart has no neutrons?

Answer Key To The Chemistry Of Life

The Princeton Review



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