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Chapter 2

Atomic Structure

2-6

- (a) Aluminum foil used for storing food weighs about 0.3 g per square cm. How many atoms of aluminum are contained in one 6.25 cm² size of foil?
- (b) Using the densities and atomic weights given in Appendix A, calculate and compare the number of atoms per cubic centimeter in (i) lead and (ii) lithium.

Solution: (a) In a 6.25 cm² sample:

$$\frac{(0.3 \text{ g}) \left(6.022 \times 10^{23} \text{ atoms/mol} \right)}{(26.981 \text{ g/mol})} = 6.7 \times 10^{21} \text{ atoms}$$

(b) (i) In lead:

$$\frac{(11.36 \text{ g/cm}^3) \left(6.022 \times 10^{23} \text{ atoms/mol} \right)}{(207.19 \text{ g/mol})} = 3.30 \times 10^{22} \text{ atoms/cm}^3$$

(ii) In lithium:

$$\frac{(0.534 \text{ g/cm}^3) \left(6.022 \times 10^{23} \text{ atoms/mol} \right)}{(6.94 \text{ g/mol})} = 4.63 \times 10^{22} \text{ atoms/cm}^3$$

Despite the different mass densities of Pb and Li, their atomic densities are approximately the same.

2-7

- (a) Using data in Appendix A, calculate the number of iron atoms in one ton (1000 kg) of iron.
- (b) Using data in Appendix A, calculate the volume in cubic centimeters occupied by one mole of boron.

Solution: (a)

$$\frac{(1000 \times 10^3) \left(6.022 \times 10^{23} \text{ atoms/mol} \right)}{(55.847 \text{ g/mol})} = 1.078 \times 10^{28} \text{ atoms}$$

Askeland Solution Manual Polymers

**Andrew J. Green, Boban Tanovic, Ian
Jones, Ann Fretwell, Peter J. Goodhew**



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illustrations in the form of microphotographs and animations A limited number of video clips are also included Page xvi

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