

# Antacid Analysis and Titration – Lab Report Assistant

## Exercise 1: Back Titration of Antacid Neutralization

**Data Table 1.** Antacid Neutralization Data

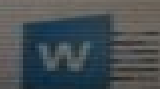
|                                |      |
|--------------------------------|------|
| Mass of Crushed Antacid (g)    |      |
| Concentration of HCl (M)       | 1M   |
| Volume HCl (mL)                | 5 mL |
| Concentration of NaOH (M)      | 1M   |
| Initial NaOH Volume (mL)       |      |
| Final NaOH Volume (mL)         |      |
| Total Volume of NaOH Used (mL) |      |

**Data Table 2.** Experimental Results

|                                                       |                |
|-------------------------------------------------------|----------------|
| HCl available for neutralization (g):                 | 0.3646 (g)     |
| Moles of NaOH required to reach stoichiometric point: | 0.0039 (moles) |
| HCl neutralized by antacid:                           | 0.222 (g)      |
| HCl neutralized per gram of antacid:                  | 0.445 (g)      |

### Questions:

- If an antacid tablet weighed 1.6 grams, how many moles of gastric acid (HCl) would it neutralize? Use the results obtained in **Data Tables 1 and 2** to explain and quantify your answer.
- If you performed this experiment with a titrant of 0.5M NaOH, would you expect your results in **Data Table 2** to change or stay the same? Explain your answer.
- The reaction that occurred when the antacid mixed with the HCl resulted in an additional product (besides a salt and water). Did you see evidence of this product? Describe the experimental evidence you witnessed that supports the formation of the additional product.



# Analysis Of Antacid Manual Answers

**John G. Koeltl**



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*Boston Journal of Chemistry and Popular Science Review*, 1871

*Laboratory Manual to Accompany The Ideas of Chemistry* Martha Orrico Visscher, 1978

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teaching basic techniques Students construct their own data tables answer conceptual questions and make predictions before performing experiments They also have the opportunity to visualize and describe molecular level activity and explain the results      *Cehmistry Textbook for College and University USA* Ibrahim Sikder,2023-06-04 Cehmistry Textbook USA

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