

100) Carbonic acid reacts as a buffer at pH 7.4. How many moles of carbonic acid are there in 1 L of blood at pH 7.4?

$$\text{pH} = 7.4 \Rightarrow \text{pK}_a = 6.35 \Rightarrow \text{pH} - \text{pK}_a = 1.05 \Rightarrow \log \frac{[\text{A}^-]}{[\text{HA}]} = 1.05 \Rightarrow \frac{[\text{A}^-]}{[\text{HA}]} = 10^{1.05} = 11.22$$

Answer = $\frac{[\text{A}^-]}{[\text{HA}]} = 11.22 \Rightarrow \text{pH} = 7.4$

101) The only intracellular buffers in body fluids are not the phosphate.

102) Describe how each of the following buffers work:

1. Phosphate buffer - acts as a buffer
in the cytoplasm
and in the extracellular fluid
in the blood
in the urine
in the sweat

103) Describe how each of the following buffers work:

1. Protein buffer - acts as a buffer
in the cytoplasm
in the extracellular fluid
in the blood
in the urine
in the sweat

104) The Henderson-Hasselbalch equation is a simplified version of the Henderson-Hasselbalch equation.

105) The Henderson-Hasselbalch equation is a simplified version of the Henderson-Hasselbalch equation.

106) The Henderson-Hasselbalch equation is a simplified version of the Henderson-Hasselbalch equation.

$$\text{pH} = 7.4 \Rightarrow \text{pK}_a = 6.35 \Rightarrow \text{pH} - \text{pK}_a = 1.05 \Rightarrow \log \frac{[\text{A}^-]}{[\text{HA}]} = 1.05 \Rightarrow \frac{[\text{A}^-]}{[\text{HA}]} = 10^{1.05} = 11.22$$



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Arshad Iqbal



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