

Hydrolysis -

The Reaction of Ions with Water

Lab Report

Purpose

The purpose of this lab is to predict whether a substance is acidic, basic, or neutral and to use the universal indicator solution to test it.

Hypothesis

Sodium Acetate	$\text{NaCH}_3\text{COO}_{(s)} \rightarrow \text{Na}^+_{(aq)} + \text{CH}_3\text{COO}^-_{(aq)}$ $\text{CH}_3\text{COO}^-_{(aq)} + \text{H}(\text{OH})_{(l)} \leftrightarrow \text{CH}_3\text{COOH}_{(aq)} + \text{OH}^-_{(aq)}$		Basic
Sodium Chloride	$\text{NaCl}_{(s)} \leftrightarrow \text{Na}^+_{(aq)} + \text{Cl}^-_{(aq)}$		Neutral
Ammonium Chloride	$\text{NH}_4\text{Cl}_{(s)} \rightarrow \text{NH}_4^+_{(aq)} + \text{Cl}^-_{(aq)}$ $\text{NH}_4^+_{(aq)} + \text{H}(\text{OH})_{(l)} \leftrightarrow \text{NH}_4(\text{OH})_{(aq)} + \text{H}^+_{(aq)}$		Acidic
Ammonium Sulfate	$(\text{NH}_4)_2\text{SO}_{4(s)} \rightarrow \text{NH}_4^+_{(aq)} + \text{SO}_4^{2-}_{(aq)}$ $\text{NH}_4^+_{(aq)} + \text{H}(\text{OH})_{(l)} \leftrightarrow \text{NH}_4(\text{OH})_{(aq)} + \text{H}^+_{(aq)}$ $\text{SO}_4^{2-}_{(aq)} + \text{H}(\text{OH})_{(l)} \leftrightarrow \text{HSO}_4^-_{(aq)} + \text{OH}^-_{(aq)}$	$K_a = 5.4 \times 10^{-10}$	Acidic
		$K_b = 8.3 \times 10^{-13}$	
Aluminum Chloride	$\text{AlCl}_{3(s)} \rightarrow \text{Al}^{3+}_{(aq)} + \text{Cl}^-_{(aq)}$ $\text{Al}^{3+}_{(aq)} + \text{H}(\text{OH})_{(l)} \leftrightarrow \text{Al}(\text{OH})_{(aq)} + \text{H}^+_{(aq)}$		Acidic
Calcium Nitrate	$\text{Ca}(\text{NO}_3)_{2(s)} \rightarrow \text{Ca}^{2+}_{(aq)} + \text{NO}_{3(aq)}$		Neutral
Iron(III) Sulfate	$\text{Fe}_2(\text{SO}_4)_{3(s)} \rightarrow \text{Fe}^{3+}_{(aq)} + \text{SO}_4^{2-}_{(aq)}$ $\text{Fe}^{3+}_{(aq)} + \text{H}(\text{OH})_{(l)} \leftrightarrow \text{Fe}(\text{OH})^{2+}_{(aq)} + \text{H}^+_{(aq)}$ $\text{Fe}^{3+}_{(aq)} + \text{H}(\text{OH})_{(l)} \leftrightarrow \text{HFe}^{2+}_{(aq)} + \text{OH}^-_{(aq)}$ $\text{SO}_4^{2-}_{(aq)} + \text{H}(\text{OH})_{(l)} \leftrightarrow \text{HSO}_4^-_{(aq)} + \text{OH}^-_{(aq)}$	$K_a = 6.0 \times 10^{-3}$	Acidic
		$K_b = 1.7 \times 10^{-12}$	
		$K_b = 8.3 \times 10^{-13}$	
Sodium Carbonate	$\text{Na}_2\text{CO}_{3(s)} \rightarrow \text{Na}^+_{(aq)} + \text{CO}_3^{2-}_{(aq)}$ $\text{CO}_3^{2-}_{(aq)} + \text{H}(\text{OH})_{(l)} \leftrightarrow \text{HCO}_3^-_{(aq)} + \text{OH}^-_{(aq)}$		Basic

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Table of Contents Answer Of Heath Chemistry Lab 20

1. Understanding the eBook Answer Of Heath Chemistry Lab 20
 - The Rise of Digital Reading Answer Of Heath Chemistry Lab 20
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer Of Heath Chemistry Lab 20
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Of Heath Chemistry Lab 20
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Of Heath Chemistry Lab 20
 - Personalized Recommendations
 - Answer Of Heath Chemistry Lab 20 User Reviews and Ratings
 - Answer Of Heath Chemistry Lab 20 and Bestseller Lists
5. Accessing Answer Of Heath Chemistry Lab 20 Free and Paid eBooks
 - Answer Of Heath Chemistry Lab 20 Public Domain eBooks
 - Answer Of Heath Chemistry Lab 20 eBook Subscription Services

- Answer Of Heath Chemistry Lab 20 Budget-Friendly Options
- 6. Navigating Answer Of Heath Chemistry Lab 20 eBook Formats
 - ePub, PDF, MOBI, and More
 - Answer Of Heath Chemistry Lab 20 Compatibility with Devices
 - Answer Of Heath Chemistry Lab 20 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answer Of Heath Chemistry Lab 20
 - Highlighting and Note-Taking Answer Of Heath Chemistry Lab 20
 - Interactive Elements Answer Of Heath Chemistry Lab 20
- 8. Staying Engaged with Answer Of Heath Chemistry Lab 20
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer Of Heath Chemistry Lab 20
- 9. Balancing eBooks and Physical Books Answer Of Heath Chemistry Lab 20
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer Of Heath Chemistry Lab 20
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answer Of Heath Chemistry Lab 20
 - Setting Reading Goals Answer Of Heath Chemistry Lab 20
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answer Of Heath Chemistry Lab 20
 - Fact-Checking eBook Content of Answer Of Heath Chemistry Lab 20
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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