

Name: _____ Course & Year: _____ Score: _____
 Schedule: _____ Date Performed: _____ Date Submitted: _____
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Activity No. 4
 WORKSHEET

CALORIMETRY AND SPECIFIC HEAT

Data, results, and calculations: Complete the table below, assume that the metal used is iron

(a) Weight of the calorimeter (inner vessel)	9.15 g
(b) Weight of Calorimeter + Water	36.70 g
(c) Weight of water, (b) - (a)	
(d) Initial temperature of water, (t) _w	28 °C
(e) Initial temperature of aluminum calorimeter, (t) _{cal}	
(f) Initial temperature of metal (t) _{metal}	98.0 °C
(g) Weight of metal sample, m _{metal}	7.5 g
(h) Final temperature of water, (t) _w	32.5 °C
(i) Final temperature of calorimeter, (t) _{cal}	
(j) Final temperature of metal	
(k) Heat absorbed by the calorimeter, q _{cal}	
(l) Heat absorbed by the water, q _w	
(m) Experimental specific heat of metal sample	
(n) Theoretical specific heat of metal sample	0.460 cal/g°C
(o) Percent Error	%

GUIDE QUESTIONS

1	What is the purpose of stirring the contents of the calorimeter before the final temperature is taken?	
2	Why is it required to tap the hot metal with a tissue paper before placing the hot metal in the calorimeter?	
3	Why should the transferring of the hot metal from the beaker to the calorimeter be done quickly?	

Answer Sheet For Calorimetry Specific Heat

G Orfield



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Table of Contents Answer Sheet For Calorimetry Specific Heat

1. Understanding the eBook Answer Sheet For Calorimetry Specific Heat
 - The Rise of Digital Reading Answer Sheet For Calorimetry Specific Heat
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer Sheet For Calorimetry Specific Heat
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Sheet For Calorimetry Specific Heat
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Sheet For Calorimetry Specific Heat

- Personalized Recommendations
 - Answer Sheet For Calorimetry Specific Heat User Reviews and Ratings
 - Answer Sheet For Calorimetry Specific Heat and Bestseller Lists
5. Accessing Answer Sheet For Calorimetry Specific Heat Free and Paid eBooks
 - Answer Sheet For Calorimetry Specific Heat Public Domain eBooks
 - Answer Sheet For Calorimetry Specific Heat eBook Subscription Services
 - Answer Sheet For Calorimetry Specific Heat Budget-Friendly Options
 6. Navigating Answer Sheet For Calorimetry Specific Heat eBook Formats
 - ePub, PDF, MOBI, and More
 - Answer Sheet For Calorimetry Specific Heat Compatibility with Devices
 - Answer Sheet For Calorimetry Specific Heat Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answer Sheet For Calorimetry Specific Heat
 - Highlighting and Note-Taking Answer Sheet For Calorimetry Specific Heat
 - Interactive Elements Answer Sheet For Calorimetry Specific Heat
 8. Staying Engaged with Answer Sheet For Calorimetry Specific Heat
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer Sheet For Calorimetry Specific Heat
 9. Balancing eBooks and Physical Books Answer Sheet For Calorimetry Specific Heat
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer Sheet For Calorimetry Specific Heat
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Answer Sheet For Calorimetry Specific Heat
 - Setting Reading Goals Answer Sheet For Calorimetry Specific Heat
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Answer Sheet For Calorimetry Specific Heat

- Fact-Checking eBook Content of Answer Sheet For Calorimetry Specific Heat
- Distinguishing Credible Sources

13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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