

Lab 10: Calculations with a Chemical Structure

Objectives:

1. To derive the structural formula of a given chemical compound without performing stoichiometric calculations.
2. To calculate the number of moles of a solid from weighing information in the lab.
3. To determine the molar mass of a solid.
4. To determine the stoichiometric amount of solid reagents that would produce a given amount of product.
5. To compare the theoretical amount to the actual amount of substance formed and calculate the percent yield.

Reagents and Apparatus:

Reagents: 0.1M Sodium Chloride Solution 25.0 mL

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Procedure:

Reagents and Calculations:

1. Calculations:

$$\text{moles of sodium chloride} = \frac{\text{volume of NaCl solution (L)}}{\text{molar mass of NaCl}} = \frac{0.1 \text{ M} \times 0.25 \text{ L}}{58.44 \text{ g/mol}} = 0.00428 \text{ mol}$$

$$\text{moles of sodium chloride} = \frac{\text{mass of NaCl}}{\text{molar mass of NaCl}} = \frac{0.25 \text{ g}}{58.44 \text{ g/mol}} = 0.00428 \text{ mol}$$

$$\text{mass of sodium chloride produced} = \text{moles of NaCl} \times \text{molar mass of NaCl} = 0.00428 \text{ mol} \times 58.44 \text{ g/mol} = 0.25 \text{ g}$$

$$\text{mass of product as found (g)} = \text{mass of NaCl} = 0.25 \text{ g}$$

2. Chemical Equations:

3. Reaction:

$$\text{NaCl} + \text{AgNO}_3 \rightarrow \text{AgCl} + \text{NaNO}_3$$

$$\text{NaCl} + \text{AgNO}_3 \rightarrow \text{AgCl} + \text{NaNO}_3$$

$$\text{moles of NaCl} = \frac{\text{mass of NaCl}}{\text{molar mass of NaCl}} = \frac{0.25 \text{ g}}{58.44 \text{ g/mol}} = 0.00428 \text{ mol}$$

Answers For Lab 6b Heath Chemistry

Tobias Bleicker



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