

Balancing redox equations

Example

Balance the equation showing the oxidation of Fe^{2+} ions to Fe^{3+} ions by dichromate ions ($\text{Cr}_2\text{O}_7^{2-}$) in an acidic medium



Step 1: Identify oxidising and reducing agents and write half reactions



$$2\text{Cr} = 12$$

$$\text{Cr} = +6$$



Iron loses one electron, i.e. it is *oxidised*. It donates one electron to dichromate, i.e. it is the *reducing agent*.

Half equation: Oxidation reaction: $\text{Fe}^{2+} \rightarrow \text{Fe}^{3+} + \text{e}^-$



Chromium gains three electrons, i.e. it is *reduced*. It gains its electrons from iron, i.e. it is the *oxidising agent*.

Half equation: Reduction reaction: $\text{Cr}_2^{6+} + 3\text{e}^- \rightarrow \text{Cr}^{3+}$

Analyzing Oxidation Reduction Reactions Answers

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Analyzing Oxidation Reduction Reactions Answers:

Limnological Analysis Robert G. Wetzel, Gene E. Likens, 2013-06-29 Limnological Analyses a classic second thoroughly updated edition consists of a series of carefully designed and tested field and laboratory exercises covering the full scope of limnology It provides the student with a solid foundation in this complex multidisciplinary field of ecology and illustrates modern experimental approaches Among the topics covered by such exercises are major physical components of lakes and streams important mineral nutrients cycling of organic matter benthic fauna primary productivity of phytoplankton quantitative methods in biota analysis diurnal changes experimental manipulation of model ecosystems effects of sewage outfall and other human activities whole ecosystem and community analyses Each exercise is preceded by an introductory section and concludes with questions for the student and a selection of suggested reading Teachers and students of limnology will value Limnological Analyses for its highly structured concise presentation Its research oriented approach encourages active participation

Volumetric Analysis: Titration methods: Oxidation-Reduction Reactions Izaak Maurits Kolthoff, Vernon Arthur Stenger, 1957 *Solutions Manual to Accompany General Chemistry with Qualitative Analysis, Second Edition* Ralph H. Petrucci, Robert K. Wismer, 1987 **Select Methods in Quantitative Analysis** Byron William Cheever, 1888 **An Introductory Course of Quantitative Chemical Analysis** Henry Paul Talbot, 1897 **Qualitative Chemical Analysis** Albert Benjamin Prescott, Otis Coe Johnson, 1892 Qualitative Chemical Analysis. A Guide in the Practical Study of Chemistry and in the Work of Analysis ... Second Edition, Revised Silas H. DOUGLAS (and PRESCOTT (Albert B.)), 1876 *Experiments and Exercises in Basic Chemistry* Steven Murov, Brian Stedjee, 2000-02-03 Lessons in Qualitative and Volumetric Chemical Analysis Charles O. Curtman, 1889 Agricultural Chemical Analysis Percy Frankland, 1883 **Russian Journal of Physical Chemistry**, 1988-09 *Volumetric Analysis* Izaak Maurits Kolthoff, 1928

Agricultural chemical analysis. Founded upon 'Leitfaden für die agricultur-chemische Analyse' von F. Krock Percy Faraday Frankland, 1883 Introduction to Drug Analysis Robert Francis Raffauf, Victor D. Warner, 1978

Fundamentals of Chemistry Laboratory Studies Frank Brescia, 1980 **American Laboratory**, 1982

Quantitative Chemical Analysis Charles Morse Allen, 1913 **Analysis of Metallurgical and Engineering Materials** Henry Wysor, 1912 **Analysis Instrumentation**, 1969 Contains Proceedings of the annual Analysis Instrumentation Symposium **A Cyclopaedia of Quantitative Chemical Analysis** Frank Humphreys Storer, 1873

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