BASIC CHEMISTRY - Part 2

Name: Answer Key
Date:
Pd:

PART 1: MIXTURES:

Fill in the blanks from the information provided on the presentation.

- A mixture is two or more substances blended together (not bonded).
- Each substance in a mixture keeps all of its chemical and physical properties.
- Mixtures, unlike compounds, can be separated into their individual substances by physical means.
- Example: sait and water can be separated by evaporation.
- A solution is a mixture of two or more substances, one dissolved in another.
- Two parts of a solution are the solute and the solvent.
- Solute substance being dissolved

Examples: sugar, salt

Solvent - substance doing the dissolving

Example: water

- A solution becomes <u>saturated</u> when the solvent <u>holds</u> all the solute that it can. It can <u>no longer</u> dissolve solute.
- Suspensions are a mixture of particles scattered throughout another. The particles do not dissolve.
 Examples: clouds dust fog. smoke, mud. blood

Reflect:

- Give another example of solutions including the solute and solvent.
 Vinegar and Calcium Carbonate, etc...
- Can solutions be separated by chemical means?
 Perhaps, but will still involve physical separation as well

PART 2: COMPOUNDS

- <u>Compounds</u> are <u>two</u> or <u>more</u> elements <u>chemically</u> joined by <u>bonds</u>.
- They are represented by chemical formulas.
- <u>Properties</u> of compounds are usually <u>different</u> than the atoms.

Example: Hydrogen(gas) + Oxygen(gas) = Water (liquid)

- Example: Iron + Oxygen = Rust
- Compounds can be <u>separated</u> into their elements only if the chemical <u>bonds</u> are <u>broken</u>. This can be done by <u>chemical</u> means (chemical <u>reaction</u>)
- A molecule is the smallest part of a compound—like water—that has all the properties of the compound.

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