

CHAPTER 10: Reactions

Reaction Energy**SECTION 1**

Write a balanced chemical equation for the following reactions in the space provided.

- For elements in their standard state, the value of ΔH_f° is _____.
- The formation and decomposition of a substance is represented by the following thermodynamic equations:



- Is the energy being added to or is it being subtracted from the decomposition?
- What is the correct sign for the enthalpy change in this decomposition reaction?

Write the correct answer in the space to the left. Show all your work in the space provided.

- _____ $\Delta H_{\text{f}}^\circ$ of a substance is 2074°C . Absorption of 4.000 J of energy will result in the decomposition of _____.
- _____ A substance has a specific heat of $0.001 \text{ J/g}^\circ\text{C}$. If a sample weighing 1.0 g is heated from 0°C to 100°C , it will absorb _____ J of energy. If the sample starts at 0°C , how much energy will it absorb if it is heated to 100°C ?
- The product is a molecular compound with a molar mass of 100 g/mol , and the substance has an enthalpy of 100 kJ/mol .
 - What is the value of ΔH for this reaction?

Answers For Reaction Energy Section 1

Ian Pickup



Answers For Reaction Energy Section 1:

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