

A) 1 B) 2 C) 3 D) 4 E) 5

55. For the endothermic reaction: $\text{CaCO}_{3(s)} \rightarrow \text{CaO}_{(s)} + \text{CO}_{2(g)}$ Le Chatelier's principle predicts that _____ will result in an increase in the number of moles of CO_2 .

- A) increasing the temperature B) decreasing the temperature
C) increasing the pressure D) removing some of the $\text{CaCO}_{3(s)}$
E) none of the above

56. The reaction, $\text{PCl}_5(\text{g}) + \text{Cl}_2(\text{g}) \rightleftharpoons \text{PCl}_6(\text{g})$ $K=5.41$ has the K shown. If a reaction mixture has the concentrations as shown, in what direction is the reaction proceeding?

Concentrations: $[\text{PCl}_5]=5.16 \times 10^{-4}\text{M}$ $[\text{Cl}_2]=1.16 \times 10^{-4}\text{M}$ $[\text{PCl}_6]=1.12 \times 10^{-4}\text{M}$

- A) Right to left B) Left to Right
C) It is at equilibrium D) Not enough information to determine

57. 1.6 mol of $\text{CH}_3\text{OH}(\text{g})$ are injected into a 4.0 L container and the following equilibrium becomes established. $2\text{H}_2(\text{g}) + \text{CO}(\text{g}) \rightleftharpoons \text{CH}_3\text{OH}(\text{g}) + 9.2 \text{ kJ}$

If at equilibrium 0.80 mol of CH_3OH is still in the container the K_{eq} must be which of the following?

- A) 0.78 B) 25 C) 5.0 D) 0.16 E) 6.25

58. Consider the following equilibrium:



At equilibrium, a 1.00 L flask contains 0.030 mol NOBr and 0.030 mol $\text{NO}_{(g)}$.

How many mol Br_2 are present?

- A) 1.9×10^4 mol B) 6.4×10^4 mol C) 3.0×10^4 mol D) 4.7×10^4 mol

59. Which will cause an increase in the numerical value of the equilibrium constant for this reaction?



- A) decreasing the temperature C) finely powdering the C(s)
B) decreasing the volume D) increasing $[\text{H}_2]$

60. What is the K_{eq} expression for



- A) $K_{\text{eq}} = [\text{H}^{+}] / [\text{Sb}^{3+}][\text{Cl}^{-}]$ B) $K_{\text{eq}} = [\text{H}^{+}][\text{SbOCl}] / [\text{Sb}^{3+}][\text{Cl}^{-}]$
C) $K_{\text{eq}} = [\text{H}^{+}]^2 / [\text{Sb}^{3+}][\text{Cl}^{-}][\text{H}_2\text{O}]$ D) $K_{\text{eq}} = [\text{H}^{+}][\text{SbOCl}] / [\text{Sb}^{3+}][\text{Cl}^{-}][\text{H}_2\text{O}]$

61. The pH of a 0.10 mol/L HCl is about 1.0 and the pH of a 0.10 mol/L H_3PO_4 is about 1.6. Which of the following best explains the difference?

- A) HCl is more ionized than H_3PO_4 in water C) H_3PO_4 is amphiprotic in water

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web andrew lee983 terms in this set 11 demand combination of quantities that someone would be willing and able to buy over a range of possible prices at a given moment