

AP[®] CHEMISTRY
2018 SCORING GUIDELINES

Question 1



Elemental sulfur can exist as molecules with the formula S_8 . The S_8 molecule is represented by the incomplete Lewis diagram above.

- (a) The diagram of S_8 shows only bonding pairs of electrons. How many lone pairs of electrons does each S atom in the molecule have?

Two	1 point is earned for the correct answer.
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- (b) Based on your answer to part (a), determine the expected value of the S–S–S bond angles in the S_8 molecule.

109.5° Acceptable range: $104^\circ \leq \text{angle} \leq 110^\circ$. (The experimentally determined angle is 107.8°)	1 point is earned for an angle in the acceptable range.
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- (c) Write the electron configuration for the S atom in its ground state.

$1s^2 2s^2 2p^6 3s^2 3p^4$ OR $[\text{Ne}] 3s^2 3p^4$	1 point is earned for a correct configuration.
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- (d) The complete photoelectron spectrum for the element chlorine is represented below. Peak X in the spectrum corresponds to the binding energy of electrons in a certain orbital of chlorine atoms. The electrons in this orbital of chlorine have a binding energy of 273 MJ/mol, while the electrons in the same orbital of sulfur atoms have a binding energy of 239 MJ/mol.

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George Ehrenhaft



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