

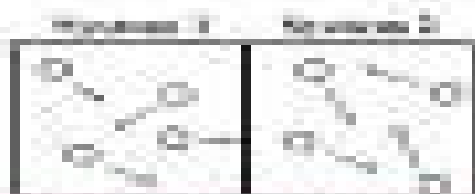
Unit 6: Thermochemistry**6.1: Heat Transfer and Thermal Equilibrium****6.2: Heat Capacity and Calorimetry****6.3: Entropy and Spontaneous Changes**

2. Use the Boltzmann-probability distribution curves below to predict temperatures of the mixtures of particles.



Temperature: is a measure of the average kinetic energy of the particles in a system. The kinetic energy for particles moving 100 m/s is 100 J . The particles in system A are moving faster and changing, resulting in a higher average kinetic energy. The Boltzmann distribution curves show that the average kinetic energy of the particles in system A is higher than the average kinetic energy of the particles in system B.

3. Sketch a curve to represent the relative distribution of the particles in the following two gases, samples of the same gas.



- a. In System 1, it is a higher or lower temperature than System 2? Justify your answer.
System 1 is at a higher temperature than System 2. Temperature is a measure of the average kinetic energy of the particles in a system. The particles in System 1 are moving faster and changing, resulting in a higher average kinetic energy. The Boltzmann distribution curves show that the average kinetic energy of the particles in System 1 is higher than the average kinetic energy of the particles in System 2.
- b. In which direction will thermal energy be transferred? Justify your answer in terms of molecular collisions.
Thermal energy will be transferred from System 1 to System 2. The particles in System 1 are moving faster and changing, resulting in a higher average kinetic energy. The Boltzmann distribution curves show that the average kinetic energy of the particles in System 1 is higher than the average kinetic energy of the particles in System 2. The particles in System 1 are moving faster and changing, resulting in a higher average kinetic energy. The Boltzmann distribution curves show that the average kinetic energy of the particles in System 1 is higher than the average kinetic energy of the particles in System 2.

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