Pre-lab Questions for Experiment #5

Rates of Chemical Reactions I: A Clock Reaction

1) Complete and balance the following reactions:

(a)
$$S_2O_8^{2-} + \Gamma \rightarrow$$

(b)
$$I_2 + S_2O_3^{2-} \rightarrow$$

2) The rate law for the reaction

$$2H_2(g) + 2NO(g) \longrightarrow N_2(g) + 2H_2O(g)$$

is rate = k[H₂][NO]². Which of the following mechanisms can be ruled out on the basis of the observed rate expression?

$$\begin{array}{c} \textit{Mechanism I} \\ \text{H}_2 + \text{NO} \longrightarrow \text{H}_2\text{O} + \text{N} & (\text{slow}) \\ \text{N} + \text{NO} \longrightarrow \text{N}_2 + \text{O} & (\text{fast}) \\ \text{O} + \text{H}_2 \longrightarrow \text{H}_2\text{O} & (\text{fast}) \\ \end{array}$$

$$\begin{array}{c} \textit{Mechanism II} \\ \text{H}_2 + 2\text{NO} \longrightarrow \text{N}_2\text{O} + \text{H}_2\text{O} & (\text{slow}) \\ \text{N}_2\text{O} + \text{H}_2 \longrightarrow \text{N}_2 + \text{H}_2\text{O} & (\text{fast}) \\ \end{array}$$

$$\begin{array}{c} \textit{Mechanism III} \\ 2\text{NO} \Longrightarrow \text{N}_2\text{O}_2 & (\text{fast equilibrium}) \\ \text{N}_2\text{O}_2 + \text{H}_2 \longrightarrow \text{N}_2\text{O} + \text{H}_2\text{O} & (\text{slow}) \\ \end{array}$$

$$\begin{array}{c} \text{N}_2\text{O}_2 + \text{H}_2 \longrightarrow \text{N}_2\text{O}_2 + \text{H}_2\text{O} & (\text{fast}) \\ \end{array}$$

3) Consider the reaction

From the following data obtained at a certain temperature, determine the order of the reaction and calculate the rate constant:

[A] (M)	[B] (M)	Rate (M/s)
1.50	1.50	3.20×10^{-1}
1.50	2.50	3.20×10^{-1}
3.00	1.50	6.40×10^{-1}

4) The following gas-phase reaction was studied at 290°C by observing the change in pressure as a function of time in a constant-volume vessel:

$$CICO_2CCI_3(g) \longrightarrow 2COCI_2(g)$$

Determine the order of the reaction and the rate constant based on the following data:

Time (s)	Partial Pressure of CICO2CCI3 (mmHg)	
0	15.76	
181	12.64	
513	8.73	
1164	4.44	

Answer Of Pre Lab Rates Of Reaction

Daniel R. Palleros

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Chemistry

Lemay, Prentice-Hall Staff, 1996

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Gilbert, Stephen F. Martin, 1994 Chemistry McGraw-Hill Staff, 2001-07 Pre-Lab Exercises for Modern

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Table of Contents Answer Of Pre Lab Rates Of Reaction

- 1. Understanding the eBook Answer Of Pre Lab Rates Of Reaction
 - The Rise of Digital Reading Answer Of Pre Lab Rates Of Reaction
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Answer Of Pre Lab Rates Of Reaction
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Of Pre Lab Rates Of Reaction
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Answer Of Pre Lab Rates Of Reaction
 - Personalized Recommendations
 - Answer Of Pre Lab Rates Of Reaction User Reviews and Ratings
 - Answer Of Pre Lab Rates Of Reaction and Bestseller Lists
- 5. Accessing Answer Of Pre Lab Rates Of Reaction Free and Paid eBooks
 - Answer Of Pre Lab Rates Of Reaction Public Domain eBooks
 - Answer Of Pre Lab Rates Of Reaction eBook Subscription Services
 - Answer Of Pre Lab Rates Of Reaction Budget-Friendly Options

- 6. Navigating Answer Of Pre Lab Rates Of Reaction eBook Formats
 - o ePub, PDF, MOBI, and More
 - Answer Of Pre Lab Rates Of Reaction Compatibility with Devices
 - Answer Of Pre Lab Rates Of Reaction Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - o Adjustable Fonts and Text Sizes of Answer Of Pre Lab Rates Of Reaction
 - Highlighting and Note-Taking Answer Of Pre Lab Rates Of Reaction
 - Interactive Elements Answer Of Pre Lab Rates Of Reaction
- 8. Staying Engaged with Answer Of Pre Lab Rates Of Reaction
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer Of Pre Lab Rates Of Reaction
- 9. Balancing eBooks and Physical Books Answer Of Pre Lab Rates Of Reaction
 - Benefits of a Digital Library
 - o Creating a Diverse Reading Collection Answer Of Pre Lab Rates Of Reaction
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answer Of Pre Lab Rates Of Reaction
 - Setting Reading Goals Answer Of Pre Lab Rates Of Reaction
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answer Of Pre Lab Rates Of Reaction
 - Fact-Checking eBook Content of Answer Of Pre Lab Rates Of Reaction
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements

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