General Chemistry II Equation Sheet

Think about how to set up the problem first, then apply the needed principles and formulas.

Phase Changes

$$\Delta T_b = iK_b m$$

 $\Delta T_f = iK_f m$
 $\pi = iMRT$
 $c = kP$
 $\frac{c_1}{P_1} = \frac{c_2}{P_2}$
 $P_{solution} = \chi_{solvent}P_{solvent}^o$
 $\Delta P = P_{solvent}^o - P_{solution} = \chi_{solute}P_{solvent}^o$

Cells

$$E = E^{\circ} - \frac{RT}{nF} \ln(Q)$$

 $E = E^{\circ} - \frac{0.0592}{n} \log(Q)$ (at $25^{\circ}C$)
 $\Delta G = -nFE_{cell}$
 $\Delta G^{\circ} = -nFE_{cell}$
 $E^{\circ}_{cell} = \frac{RT}{nF} \ln(K)$
 $E^{\circ}_{cell} = \frac{0.0592V}{n} \log(K)$ (at $25^{\circ}C$)
USE ONE:

$$E_{cell}^{\circ} = E_{cathode(red.)}^{\circ} - E_{anode(red.)}^{\circ}$$

 $E_{cell}^{\circ} = E_{cathode(red.)}^{\circ} + E_{anode(ox.)}^{\circ}$

Thermodynamics

$$\Delta G = \Delta H - T\Delta S$$

 $\Delta S_{surr} = \frac{-\Delta H_{eys}}{I}$
 $\Delta S_{univ} = \Delta S_{sys} + \Delta S_{surr}$
 $\Delta S_{rsw}^{\circ} = \Sigma S_{products}^{\circ} + \Sigma S_{reactants}^{\circ}$
 $\Delta G_{csm}^{\circ} = \Sigma \Delta G_{products}^{\circ} + \Sigma \Delta G_{reactants}^{\circ}$

Concentrations

$$\begin{array}{ll} \text{mobality} = \frac{modes_{column}}{modes_{column}} \\ \% \text{ by mass} = \frac{modes_{column}}{modes_{column}} \times 100\% \\ \text{parts per million} = \frac{modes_{column}}{modes_{column}} \times 10^6 \\ \text{parts per billion} = \frac{modes_{column}}{modes_{column}} \times 10^9 \end{array}$$

Acids and Bases

$$K_w = [H_3O^+][OH^-] = K_a \times K_b$$

 $pOH = -\log[OH^-]$
 $[OH^-] = 10^{-pOH}$
 $pH + pOH = pK_a + pK_b = 14.00 \text{ (at } 25^{\circ}C)$
 $pH = pK_a + \log \frac{[A^-]}{[HA]}$
 $pOH = pK_b + \log \frac{[BH^+]}{[B]}$
% ionization $\frac{[H_3O^+]_{eq}}{[HA]_a} \times 100\%$
 $pK_a = -\log(K_a)$
 $pK_b = -\log(K_b)$
 $K_a = 10^{-pK_b}$

Equilibria

$$Q_c = \frac{|C|^c |D|^d}{|A|^a |B|^b}$$
 $K_c = \frac{|C|^c |D|^d}{|A|^a |B|^b}$
 $Q_p = \frac{P_A^d P_C^d}{P_A^a P_B^a}$
 $K_p = \frac{P_A^d P_C^d}{P_A^a P_B^a}$
 $K_p = K_c [R \times T]^{\Delta_m}$
 $\Delta G = \Delta G^o + RT \ln(Q)$
 $\Delta G^o = -RT \ln(K)$

rate = $k[A]^a [B]^a$
 $k = Ae^{-\frac{E_a}{M^2}}$
 $\ln(\frac{k_B}{k_B}) = \frac{E_a}{R}(\frac{1}{T_B} - \frac{1}{T_1})$

half life = $\frac{\ln(2)}{k}$ (1st order reactions only)

 $[A]_d = -kt + [A]_o$
 $\ln([A]_d) = -kt + \ln([A]_o)$
 $\frac{1}{|A|_L} = kt + \frac{1}{|A|_o}$



Answer Sheet General Chemistry Equations

Stanley Wallace Morse

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- 1. Understanding the eBook Answer Sheet General Chemistry Equations
 - The Rise of Digital Reading Answer Sheet General Chemistry Equations
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Answer Sheet General Chemistry Equations
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Sheet General Chemistry Equations
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Answer Sheet General Chemistry Equations
 - Personalized Recommendations
 - Answer Sheet General Chemistry Equations User Reviews and Ratings
 - Answer Sheet General Chemistry Equations and Bestseller Lists
- 5. Accessing Answer Sheet General Chemistry Equations Free and Paid eBooks
 - Answer Sheet General Chemistry Equations Public Domain eBooks
 - Answer Sheet General Chemistry Equations eBook Subscription Services
 - Answer Sheet General Chemistry Equations Budget-Friendly Options
- 6. Navigating Answer Sheet General Chemistry Equations eBook Formats

- o ePub, PDF, MOBI, and More
- Answer Sheet General Chemistry Equations Compatibility with Devices
- Answer Sheet General Chemistry Equations Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - o Adjustable Fonts and Text Sizes of Answer Sheet General Chemistry Equations
 - Highlighting and Note-Taking Answer Sheet General Chemistry Equations
 - Interactive Elements Answer Sheet General Chemistry Equations
- 8. Staying Engaged with Answer Sheet General Chemistry Equations
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - \circ Following Authors and Publishers Answer Sheet General Chemistry Equations
- 9. Balancing eBooks and Physical Books Answer Sheet General Chemistry Equations
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer Sheet General Chemistry Equations
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answer Sheet General Chemistry Equations
 - \circ Setting Reading Goals Answer Sheet General Chemistry Equations
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answer Sheet General Chemistry Equations
 - Fact-Checking eBook Content of Answer Sheet General Chemistry Equations
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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