

6. Given the following half-reaction reduction of a galvanic cell:



- a. Identify the anode and cathode half-reactions:



Note: This problem is an iron electrode with iron ions only in solution.

- b. Write the overall balanced cell reaction:



6. Write the cell notation for a galvanic cell that uses the reactions below:



7. A galvanic cell is constructed based on the following reaction:



- a. Calculate the standard free energy change (ΔG°) for this reaction at 298 K.



$$= -2(96485 \text{ C/mol})E^{\circ}_{\text{cell}}$$

$$= -2(96485 \text{ C/mol})E^{\circ}_{\text{cell}}$$



$$= -2(96485 \text{ C/mol})(0.77 \text{ V} - 0.44 \text{ V})$$

$$= -2(96485 \text{ C/mol})(0.33 \text{ V})$$

$$= -2(96485 \text{ C/mol})(0.33 \text{ V})$$

- b. Is this reaction spontaneous?

Yes, this reaction is spontaneous at standard state conditions, as indicated by the negative sign of ΔG° .

Answer Keys To Problem Of The Month

**Jerry J. Weygandt, Paul D.
Kimmel, Donald E. Kieso**



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