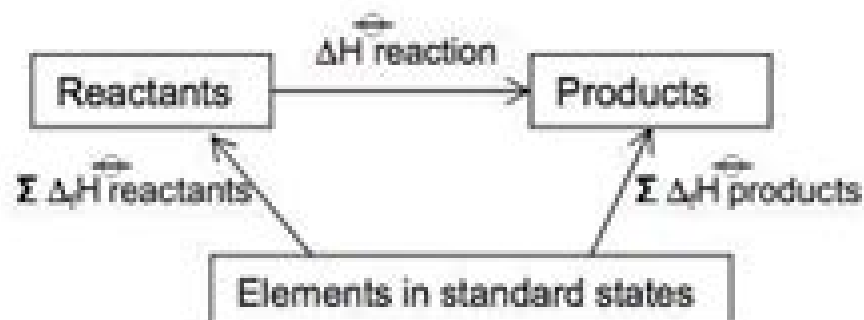


Using Hess's law to determine enthalpy changes from enthalpy changes of formation.

$$\Delta H_{\text{reaction}} = \sum \Delta_f H_{\text{products}} - \sum \Delta_f H_{\text{reactants}}$$



Example 6. What is the enthalpy change for this reaction ?



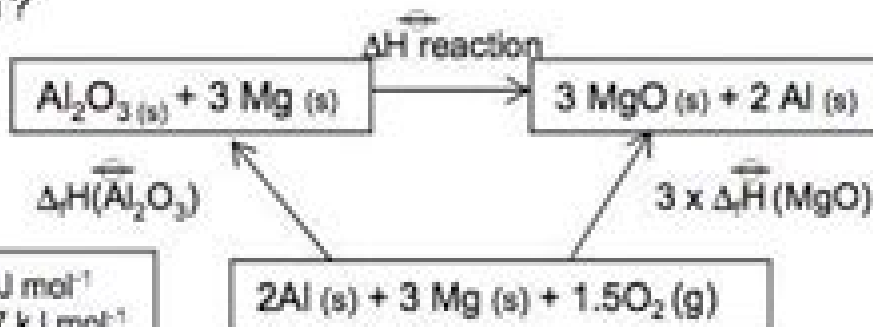
$$\Delta H = \sum \Delta_f H_{\text{products}} - \sum \Delta_f H_{\text{reactants}}$$

$$\Delta H = 3 \times \Delta_f H (\text{MgO}) - \Delta_f H (\text{Al}_2\text{O}_3)$$

$$\begin{aligned} \Delta H &= (3 \times -601.7) - -1675.7 \\ &= -129.4 \text{ kJ mol}^{-1} \end{aligned}$$

Remember
elements have
 $\Delta_f H = 0$

$$\begin{aligned} \Delta_f H (\text{MgO}) &= -601.7 \text{ kJ mol}^{-1} \\ \Delta_f H (\text{Al}_2\text{O}_3) &= -1675.7 \text{ kJ mol}^{-1} \end{aligned}$$



Example 7. Using the following data to calculate the heat of combustion of propene

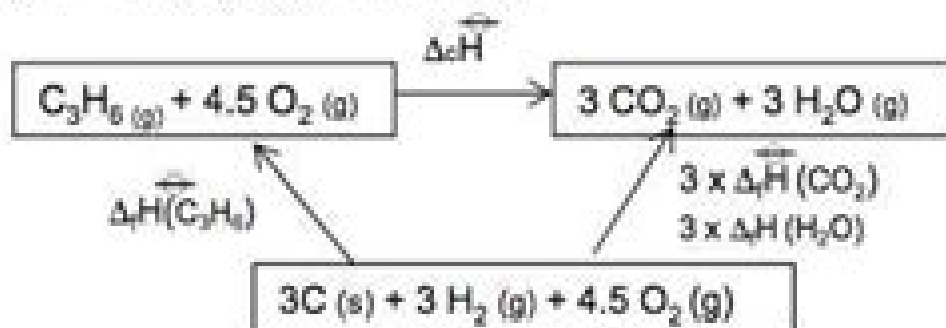
$$\Delta_f H, \text{C}_3\text{H}_6(\text{g}) = +20 \text{ kJ mol}^{-1} \quad \Delta_f H, \text{CO}_2(\text{g}) = -394 \text{ kJ mol}^{-1} \quad \Delta_f H, \text{H}_2\text{O}(\text{g}) = -242 \text{ kJ mol}^{-1}$$



$$\Delta_c H = \sum \Delta_f H_{\text{products}} - \sum \Delta_f H_{\text{reactants}}$$

$$\Delta_c H = [3 \times \Delta_f H (\text{CO}_2) + 3 \times \Delta_f H (\text{H}_2\text{O})] - \Delta_f H (\text{C}_3\text{H}_6)$$

$$\begin{aligned} \Delta_c H &= [(3 \times -394) + (3 \times -242)] - 20 \\ &= -1928 \text{ kJ mol}^{-1} \end{aligned}$$



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undergraduate course or not Thus its readership extends beyond conventional technical undergraduates to practicing engineers and also to the interested lay person who seeks to understand the discourse that surrounds the choice of particular technological solutions to current and future energy and material production problems *Thermochemistry and Thermodynamics* Henry Alistair Skinner, 1975 **Hess's Law** Lifeliqe, 2019 This lesson plan covers two indirect methods of determining enthalpy change for most chemical reactions using Hess s law

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