

1. The table below shows the number of valence electrons in atoms of elements **W**, **X**, and **Y**.

Element	W	X	Y
No. of valence electrons	1	5	2

- a) Elements **W** and **Y** cannot combine to form a compound. Explain (1 mark)

They both react by losing electrons

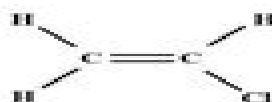
- b) The carbonate of **W** was heated in a test tube, using a Bunsen burner flame. State and explain the observations made. (2 marks)

No production of a colourless gas that forms a white precipitate in calcium hydroxide solution. Carbonates of group 1 elements do not decompose on heating.

2. One mole of ethyne reacted with one mole of hydrogen chloride gas:

- a) Draw the structure and give the name of the compound formed (2 marks)

Structure



Name:

1-chloroethene

- b) Excess hydrogen was bubbled through the product in the presence of nickel catalyst. Three drops of acidified potassium manganate (VII) were then added to the product of reaction. State the observation made when this oxidising agent was added. (1 mark)

Purple acidified potassium manganate (VII) remains purple

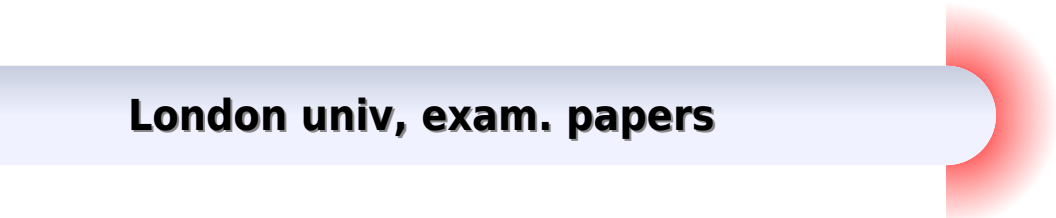
3. A certain carbonate has the formula **WCO₃**. If 1g of the carbonate reacts completely with 20cm³ of 1M hydrochloric acid, calculate the relative atomic mass of **W** (3 marks)

(O = 16, C = 12)

Moles HCl	Moles WCO ₃	R.F.M. WCO ₃
$ \frac{20 \times 1}{1000} \\ = 0.02 \text{ moles} $	$ \begin{array}{l} \text{HCl} : \text{WCO}_3 \\ 2 : 1 \\ = \frac{0.02 \times 1}{2} \\ = 0.01 \text{ moles} \end{array} $	$ \begin{array}{l} 0.01 \text{ moles} \equiv 1\text{g} \\ 1 \text{ mole} \equiv \frac{1 \times 1}{0.01} \\ = 100 \\ W + 12 + 3(16) = 100 \end{array} $

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Table of Contents Answers For Chemisty Paper

1. Understanding the eBook Answers For Chemisty Paper
 - The Rise of Digital Reading Answers For Chemisty Paper
 - Advantages of eBooks Over Traditional Books
2. Identifying Answers For Chemisty Paper
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answers For Chemisty Paper
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answers For Chemisty Paper
 - Personalized Recommendations
 - Answers For Chemisty Paper User Reviews and Ratings

- Answers For Chemistry Paper and Bestseller Lists
- 5. Accessing Answers For Chemistry Paper Free and Paid eBooks
 - Answers For Chemistry Paper Public Domain eBooks
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 - Answers For Chemistry Paper Budget-Friendly Options
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 - Answers For Chemistry Paper Enhanced eBook Features
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 - Highlighting and Note-Taking Answers For Chemistry Paper
 - Interactive Elements Answers For Chemistry Paper
- 8. Staying Engaged with Answers For Chemistry Paper
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 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answers For Chemistry Paper
- 9. Balancing eBooks and Physical Books Answers For Chemistry Paper
 - Benefits of a Digital Library
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- 10. Overcoming Reading Challenges
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 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answers For Chemistry Paper
 - Setting Reading Goals Answers For Chemistry Paper
 - Carving Out Dedicated Reading Time
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 - Fact-Checking eBook Content of Answers For Chemistry Paper
 - Distinguishing Credible Sources

-
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 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
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