STOICHIOMETRY

 Calculate the number of moles of NaOH that are needed to react with 500 g of H₂SO₄ according to the following equation:

$$H_2SO_4 + 2 NaOH \rightarrow Na_2SO_4 + 2 H_2O$$

2. Calculate the mass of NH_3 that can be produced from the reaction of $125\,g$ of NCl_3 according to the following equation:

3. Identify the limiting reactant and determine the mass of CO_2 that can be produced from the reaction of 25 g of C_3H_8 with 75 g of O_2 according to the following equation:

$$C_3H_8 + 5O_2 \rightarrow 3 CO_2 + 4 H_2O$$

4. How many grams of SO_2 are produced when 152 g of CS_2 reacts with 48.0 g of O_2 according to the following equation:

$$CS_2 + 3O_2 \rightarrow CO_2 + 2SO_2$$

Answers For The Chemistry Lab Of Stoichiometry

Siddappa N.Byrareddy

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the experimental nature of general chemistry Experiments are written to address students of various academic backgrounds and differing interests and abilities in chemistry While most experiments can be conducted in a single three hour period some have been designed to be completed over an extended time to illustrate that chemical systems do not work at an arbitrary schedule Suggestions are provided for combining experiments of shorter length and similar pedagogy

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