

**Lab 10: Calculations with a Chemical Structure**

**Objectives:**

1. To derive the empirical formula of a pure substance with percent composition (molecular formula) information.
2. To calculate the number of moles of a solid from weighing information on the balance.
3. To determine the molar mass of a solid.
4. To determine the theoretical amount of product that would be produced.
5. To compare the theoretical amount to the actual amount of substance formed and calculate the percent yield.

**Reaction Information:**

Reaction of Sodium Chloride Solution: 25.0 mL

Solution of Calcium Chloride Solution: 25.0 mL

Approx. Dry Filter Paper: 0.00 g

Approx. Filter Paper + Dry Solid: 0.00 g

Describe what happens when mixing the two solutions:

Initially the solution turned milky and cloudy. After stirring with a bar magnet, the solution became milky and the bottom portion turned black in part.

**Results:**

**Observations and Calculations:**

**(a) Calculations:**

$$\text{moles of sodium chloride} = 0.5 \text{ mol/L NaCl} \times \frac{0.025 \text{ L}}{1000 \text{ mL}} = 0.0125 \text{ mol NaCl}$$

$$\text{moles of calcium chloride} = 0.0125 \text{ mol NaCl} \times \frac{1 \text{ mol CaCl}_2}{2 \text{ mol NaCl}} = 0.00625 \text{ mol CaCl}_2$$

$$\text{mass of calcium chloride produced} = 0.00625 \text{ mol CaCl}_2 \times 110.98 \text{ g/mol} = 0.694 \text{ g}$$

$$\text{mass of product as found (produced)} = 0.00625 \text{ mol CaCl}_2 \times \frac{1 \text{ mol}}{1 \text{ mol}} = 0.00625 \text{ mol CaCl}_2$$

$$\text{Chemical equation: } \text{Na}_2\text{CO}_3 + \text{CaCl}_2 \rightarrow \text{CaCO}_3 + 2 \text{NaCl}$$

**(b) Percent Yield:**

$$\text{Theoretical yield} = 0.00625 \text{ mol CaCl}_2 \times 110.98 \text{ g/mol} = 0.694 \text{ g}$$

$$\text{Actual yield} = 0.00625 \text{ mol CaCl}_2 \times 110.98 \text{ g/mol} = 0.694 \text{ g}$$

$$\text{Percent yield} = \frac{\text{Actual yield}}{\text{Theoretical yield}} \times 100\% = \frac{0.694 \text{ g}}{0.694 \text{ g}} \times 100\% = 100\%$$

# Answers For Lab 6b Heath Chemistry

**Jianjun Gao**



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