

11.1

DESCRIBING CHEMICAL REACTIONS

Section Review

Objectives

- 1. Represent the processes of a reaction.
- 2. Explain how to write a balanced equation.
- 3. List the steps for writing a complete ionic equation.

Vocabulary

- 1. chemical equation
- 2. balanced equation
- 3. reactant
- 4. coefficient
- 5. product

Part A Completion

Write the coefficient for each in the balanced equation, and balance the equation by changing the coefficient for each substance. Write balanced ionic equations for each.

1. A chemical equation is always represented by a balanced _____ 2. _____
 The chemical formula for a chemical compound is _____ 3. _____ The mass
 balance between reactants and products in the _____ 4. _____ equation is called the
 law of conservation of _____ 5. _____ A chemical equation is written with an
 arrow pointing to the right, and gives _____ 6. _____ as reactants and
 products. The law of conservation of mass states that _____ 7. _____ mass is neither lost
 nor gained. The equation must always balance for charge. The _____ 8. _____ of
 the equation must be balanced.

9. A chemical equation is written with the products of a reaction as a
 product. The equation is _____ 10. _____ that a solid _____ 11. _____ and an aqueous
 _____ 12. _____ substance dissociates into ions. _____ 13. _____ of _____ 14. _____
 must be present to show a complete ionic equation. _____ 15. _____ the
 equation.

- 1. equation
- 2. formula
- 3. product
- 4. mass
- 5. conservation
- 6. reactants
- 7. law of conservation
- 8. charge
- 9. solid
- 10. balanced
- 11. precipitate
- 12. aqueous
- 13. ions
- 14. balanced
- 15. complete

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Christian Drosten



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