ACIDS, BASES, & THE pH SCALE

Name: Answer Key
Date:
Pd:

PART 1: ACIDS AND BASES

Comparing Acids and Bases	
Acids	Bases (Alkalines)
 Sour taste, leaves burning sensation 	 Sharp <u>bitter</u> taste, feels <u>slippery</u> to touch
 Increase hydrogen ions (H+) in solution 	 Decrease hydrogen ions (H+) in solution (Increase hydroxide ions (OH-) in solution)
• Example: HCI H' + CI	Example: NaOH Na' + OH
 Turns <u>litmus</u> (pH) paper to <u>red</u> 	 Turns litmus (pH) paper to blue

PART 2: THE pH SCALE

- The pH Scale is a measurement of how acidic or how basic a solution is.
- The scale goes from 0 to 14.
- . Compounds are acidic if they have a pH lower than Z.
- Compounds with a pH <u>higher</u> than <u>7</u> are <u>basic</u> or <u>alkaline</u>.
- Compounds with a pH of 7, are neutral.

0 7 8 9 10 11 12 13

Increased concentration of H ions
Decreased concentration of OH ions

Increased concentration of OH ions

pH below 7 = acids

pH at $7 = \frac{neutral}{pure water}$

pH above 7 = <u>bases</u>

Decreased concentration of H' ions

PART 3: BUFFERS

- A <u>buffer</u> is a <u>weak acid</u> or <u>base</u> that <u>prevents</u> sudden <u>changes</u> in pH.
 - Maintains homeostasis
- A buffer works by binding to H+ ions releasing H+ ions when the solution pH changes.
- Buffers bind to H+ ions when the concentration of H= ions in solution increases.
- NaOH Na + OH (The OH- will bind to H+ ions to form H;O neutralization)
- Buffers <u>release</u> H+ ions when the <u>concentration</u> of H= ions in solution <u>decreases</u>.
 HCI H' + CI

20 Acids And Bases Answer Key

Frank R. Foulkes

20 Acids And Bases Answer Key:

General Chemistry Donald A. McQuarrie, Stanley Gill, 2011-06-15 This Fourth Edition of McQuarrie's classic text offers a thorough revision and a quantum leap forward from the previous edition Taking an atoms first approach it promises to be another ground breaking text in the tradition of McQuarrie's many previous works. This outstanding new text available in a soft cover edition offers professors a fresh choice and outstanding value **Physical Chemistry for Engineering and Applied Sciences** Frank R. Foulkes, 2012-09-12 Physical Chemistry for Engineering and Applied Sciences is the product of over 30 years of teaching first year Physical Chemistry as part of the Faculty of Applied Science and Engineering at the University of Toronto Designed to be as rigorous as compatible with a first year student's ability to understand the text presents detailed step by step derivations of the equations that permit the student to follow the underlying logic and of equal importance to appreciate any simplifying assumptions made or mathematical tricks employed In addition to the 600 exercises and end of chapter problems the text is rich in worked non trivial examples many of which are designed to be inspiring and thought provoking Step by step derivation of all equations enables the student to smoothly follow the derivation by sight and can be understood relatively easily by students with moderate skills and backgrounds in mathematics Clear and accessible Physical Chemistry for Engineering and Applied Sciences includes The answers to all of the 112 worked examples 99 exercises following many of the worked examples and 496 end of chapter problems Topics not normally seen in introductory physical chemistry textbooks ionic reaction rates activities and activity coefficients or not regularly explained in much detail electrochemistry chemical kinetics with an eye on industrial applications Special appendices that provide detailed explanations of basic integration and natural logarithms for students lacking a background in integral calculus An in depth chapter on electrochemistry in which activities and activity coefficients are used extensively as required for accurate calculations **Chemistry (Teacher Guide)** Dr. Dennis Englin, 2018-02-26 This book was created to help teachers as they instruct students through the Master's Class Chemistry course by Master Books The teacher is one who guides students through the subject matter helps each student stay on schedule and be organized and is their source of accountability along the way With that in mind this guide provides additional help through the laboratory exercises as well as lessons guizzes and examinations that are provided along with the answers The lessons in this study emphasize working through procedures and problem solving by learning patterns The vocabulary is kept at the essential level Practice exercises are given with their answers so that the patterns can be used in problem solving These lessons and laboratory exercises are the result of over 30 years of teaching home school high school students and then working with them as they proceed through college Guided labs are provided to enhance instruction of weekly lessons There are many principles and truths given to us in Scripture by the God that created the universe and all of the laws by which it functions It is important to see the hand of God and His principles and wisdom as it plays out in chemistry This course integrates what God has told us in the context of this study

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