



Suggested answers to in-text activities and unit-end exercises

Topic 6 Unit 24

In-text activities

Checkpoint (page 38)

- 1 a) P–H non-polar
 $\delta^- \quad \delta^+$
 b) H–F
 $\delta^- \quad \delta^+$
 c) O–N
 $\delta^- \quad \delta^+$
 d) F–F non-polar
 $\delta^- \quad \delta^+$
 e) N–H
 $\delta^- \quad \delta^+$
 f) C–I non-polar

- 2 In NH_3 , the bond pairs of electrons are attracted towards the nitrogen atom to a greater extent as nitrogen is more electronegative than hydrogen.

The bond pairs of electrons repel each other to a greater extent and thus the H–N–H bond angle is greater.

In NF_3 , the bond pairs of electrons are closer to the fluorine atom as fluorine is more electronegative than nitrogen.

The bond pairs of electrons repel each other to a less extent and thus the F–N–F bond angle is smaller.

Checkpoint (page 44)

Molecular formula	Electron diagram	Shape of molecule	Polar bond	Polar molecule?
H_2O	$\begin{array}{c} \text{H} \times \times \text{O} \times \times \\ \times \times \times \times \\ \text{H} \end{array}$	V-shaped	O–H	yes
BeH_2	$\text{H} \times \times \text{Be} \times \times \text{H}$	linear	Be–H	no
CH_2F_2	$\begin{array}{c} \text{H} \\ \times \times \times \times \\ \text{H} \times \times \text{C} \times \times \text{F} \times \times \\ \times \times \times \times \\ \times \times \text{F} \times \times \\ \times \times \end{array}$	tetrahedral	C–F	yes
Cl_2O	$\begin{array}{c} \times \times \times \times \\ \times \times \text{Cl} \times \times \text{O} \times \times \\ \times \times \times \times \\ \times \times \text{Cl} \times \times \\ \times \times \end{array}$	V-shaped	Cl–O	yes

21st Century Chemistry Supplementary Exercises Answer

James Osler Bailey, Thomas B. Stroup



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Table of Contents 21st Century Chemistry Supplementary Exercises Answer

1. Understanding the eBook 21st Century Chemistry Supplementary Exercises Answer
 - The Rise of Digital Reading 21st Century Chemistry Supplementary Exercises Answer
 - Advantages of eBooks Over Traditional Books
2. Identifying 21st Century Chemistry Supplementary Exercises Answer
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an 21st Century Chemistry Supplementary Exercises Answer
 - User-Friendly Interface
4. Exploring eBook Recommendations from 21st Century Chemistry Supplementary Exercises Answer
 - Personalized Recommendations
 - 21st Century Chemistry Supplementary Exercises Answer User Reviews and Ratings
 - 21st Century Chemistry Supplementary Exercises Answer and Bestseller Lists
5. Accessing 21st Century Chemistry Supplementary Exercises Answer Free and Paid eBooks
 - 21st Century Chemistry Supplementary Exercises Answer Public Domain eBooks
 - 21st Century Chemistry Supplementary Exercises Answer eBook Subscription Services
 - 21st Century Chemistry Supplementary Exercises Answer Budget-Friendly Options
6. Navigating 21st Century Chemistry Supplementary Exercises Answer eBook Formats
 - ePub, PDF, MOBI, and More
 - 21st Century Chemistry Supplementary Exercises Answer Compatibility with Devices
 - 21st Century Chemistry Supplementary Exercises Answer Enhanced eBook Features
7. Enhancing Your Reading Experience

- Adjustable Fonts and Text Sizes of 21st Century Chemistry Supplementary Exercises Answer
 - Highlighting and Note-Taking 21st Century Chemistry Supplementary Exercises Answer
 - Interactive Elements 21st Century Chemistry Supplementary Exercises Answer
8. Staying Engaged with 21st Century Chemistry Supplementary Exercises Answer
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers 21st Century Chemistry Supplementary Exercises Answer
 9. Balancing eBooks and Physical Books 21st Century Chemistry Supplementary Exercises Answer
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection 21st Century Chemistry Supplementary Exercises Answer
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine 21st Century Chemistry Supplementary Exercises Answer
 - Setting Reading Goals 21st Century Chemistry Supplementary Exercises Answer
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of 21st Century Chemistry Supplementary Exercises Answer
 - Fact-Checking eBook Content of 21st Century Chemistry Supplementary Exercises Answer
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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