

AP[®] CHEMISTRY
2004 SCORING GUIDELINES

Question 1

Answer the following questions relating to the solubilities of two silver compounds, Ag_2CrO_4 and Ag_3PO_4 .

Silver chromate dissociates in water according to the equation shown below.



(a) Write the equilibrium-constant expression for the dissolving of $\text{Ag}_2\text{CrO}_4(s)$.

$$K_{sp} = [\text{Ag}^+]^2[\text{CrO}_4^{2-}]$$

1 point for correct expression

(b) Calculate the concentration, in mol L^{-1} , of $\text{Ag}^+(aq)$ in a saturated solution of Ag_2CrO_4 at 25°C .



I	—	0	0
C	—	+2x	+x
F	—	0 + 2x	0 + x

1 point for correct stoichiometry of $\text{Ag}^+(aq)$ and $\text{CrO}_4^{2-}(aq)$

$$2.6 \times 10^{-12} = [\text{Ag}^+]^2[\text{CrO}_4^{2-}]$$

$$2.6 \times 10^{-12} = [2x]^2[x] = 4x^3$$

$$8.7 \times 10^{-5} = x = [\text{CrO}_4^{2-}]$$

$$[\text{Ag}^+] = 2x = 2 \times (8.7 \times 10^{-5} M) = 1.7 \times 10^{-4} M$$

1 point for substituting and calculating $[\text{Ag}^+]$

(c) Calculate the maximum mass, in grams, of Ag_2CrO_4 that can dissolve in 100. mL of water at 25°C .

$$\frac{8.7 \times 10^{-5} \text{ mole Ag}_2\text{CrO}_4}{1 \text{ L}} \times \frac{331.7 \text{ g}}{1 \text{ mole Ag}_2\text{CrO}_4} = \frac{0.029 \text{ g Ag}_2\text{CrO}_4}{1 \text{ L}}$$

1 point for molar mass of Ag_2CrO_4

$$\frac{0.029 \text{ g Ag}_2\text{CrO}_4}{1 \text{ L}} \times 0.100 \text{ L} = 0.0029 \text{ g Ag}_2\text{CrO}_4$$

1 point for mass of Ag_2CrO_4 in 100 mL

Scoring Guidelines Ap Chemistry

Mina Lebitz



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