

Name: Key

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Extra Practice with Inscribed and Central Angles section 10.4

Directions: Find each angle measure, arc measure, or value of the variable.

1. $m\angle XNM = \underline{40^\circ}$

$m\widehat{YX} = \underline{75^\circ}$



2. $m\widehat{TR} = \underline{150^\circ}$

$m\widehat{SR} = \underline{96^\circ}$



3. $m\widehat{WXY} = \underline{180^\circ}$

$m\angle WYX = \underline{50^\circ}$



4. $m\angle J = \underline{46^\circ}$

$m\angle K = \underline{90^\circ}$

$m\angle L = \underline{44^\circ}$



5. $m\widehat{BYD} = \underline{160^\circ}$

$m\angle YBC = \underline{88^\circ}$



6. $m\angle ECD = \underline{53^\circ}$



7. $x = \underline{66^\circ}$

$y = \underline{66^\circ}$

$z = \underline{26^\circ}$

$w = \underline{70^\circ}$



8. $m\angle F = \underline{115^\circ}$

$m\widehat{FH} = \underline{122^\circ}$



9. $x = \underline{3}$

$m\widehat{FGH} = \underline{238^\circ}$



$21x - 2 + 38x + 5 = 180$

$59x + 3 = 180$

$59x = 177$

$x = 3$

10. $m\angle W = \underline{22^\circ}$

$m\angle Y = \underline{52^\circ}$

$m\angle ZVY = \underline{74^\circ}$



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Disha Experts

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the 'Disha Experts' text.

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web we can balance it for charge by adding two electrons to the right side of the equation so that the net charge on each side is 0 oxidation $\text{Ni} \rightarrow \text{Ni}^{2+} + 2\text{e}^-$ now that the oxidation half reaction is balanced it tells us that two electrons are produced for every atom of nickel oxidized but where do those electrons go

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web in the redox equation below that occurs in our stomach which of the following pairs identifies the reducing and oxidizing agents respectively $\text{C}_6\text{H}_8\text{O}_6 + 2\text{HNO}_2 + 2\text{C}_6\text{H}_6\text{O}_6 + 2\text{H}_2\text{O} \rightarrow 2\text{HNO}$ choose 1 answer $\text{C}_6\text{H}_8\text{O}_6$ and NO_2 a $\text{C}_6\text{H}_8\text{O}_6$ and NO_2 b NO_2 and $\text{C}_6\text{H}_8\text{O}_6$ c $\text{C}_6\text{H}_8\text{O}_6$ and H_2O d H_2O and $\text{C}_6\text{H}_8\text{O}_6$

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