

IONIC & COVALENT BONDS

Name:

Date:

Period:

An _____ is the attraction between two oppositely charged ions. The chemical bond formed when two atoms share electrons is called a _____. Covalent bonds usually form between atoms of _____ and ionic bonds usually form when a _____ combines with a _____.

Activity A: How Do Ions Form?

Procedure:

1. Place three pairs of checkers (three red and three black) on your desk. The red represent electrons and the black represent protons.
2. Place nine pairs of checkers (nine red and nine black) in a separate group on your desk.
3. Move a red checker from the smaller group to the larger group.
4. Count the number of positive charges (protons) and negative charges (electrons) in each group.
5. Now sort the checkers into a group of four pairs and a group of eight pairs. Repeat Steps 3 and 4, this time moving two red checkers from the smaller group to the larger group.

Think it over:

1. What was the total charge on each group before you moved the red checkers (electrons)?
2. What was the charge on each group after you moved the checkers?
3. Based on this activity what do you think happens to the charge on an atom when it loses electrons?
4. Based on this activity what do you think happens to the charge on an atom when it gains electrons?

Activity B: Interpreting Data About Ions

Procedure:

1. Use the periodic table and the chart to the right to identify the charges of the ions in each ionic compound listed below.
2. Write the formulas for each compound.
 - a. sodium bromide;
 - b. lithium oxide;
 - c. magnesium sulfide;
 - d. aluminum fluoride;
 - e. potassium nitrate;
 - f. ammonium chloride

Ions and Their Charges		
Name	Charge	Symbol or Formula
Lithium	1+	Li^+
Sodium	1+	Na^+
Potassium	1+	K^+
Ammonium	1+	NH_4^+
Calcium	2+	Ca^{2+}
Magnesium	2+	Mg^{2+}
Aluminum	3+	Al^{3+}
Fluoride	1-	F^-
Chloride	1-	Cl^-
Iodide	1-	I^-
Bicarbonate	1-	HCO_3^-
Nitrate	1-	NO_3^-
Oxide	2-	O^{2-}
Sulfide	2-	S^{2-}
Carbonate	2-	CO_3^{2-}
Sulfate	2-	SO_4^{2-}
Phosphate	3-	PO_4^{3-}

Think it over:

1. How do you know how many of each ion to write in the formula?

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1. Understanding the eBook Answer Key Ionic Bonds Review And Reinfore
 - The Rise of Digital Reading Answer Key Ionic Bonds Review And Reinfore
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer Key Ionic Bonds Review And Reinfore
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Key Ionic Bonds Review And Reinfore
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Key Ionic Bonds Review And Reinfore
 - Personalized Recommendations
 - Answer Key Ionic Bonds Review And Reinfore User Reviews and Ratings
 - Answer Key Ionic Bonds Review And Reinfore and Bestseller Lists
5. Accessing Answer Key Ionic Bonds Review And Reinfore Free and Paid eBooks
 - Answer Key Ionic Bonds Review And Reinfore Public Domain eBooks
 - Answer Key Ionic Bonds Review And Reinfore eBook Subscription Services
 - Answer Key Ionic Bonds Review And Reinfore Budget-Friendly Options
6. Navigating Answer Key Ionic Bonds Review And Reinfore eBook Formats

- ePub, PDF, MOBI, and More
- Answer Key Ionic Bonds Review And Reinfore Compatibility with Devices
- Answer Key Ionic Bonds Review And Reinfore Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answer Key Ionic Bonds Review And Reinfore
 - Highlighting and Note-Taking Answer Key Ionic Bonds Review And Reinfore
 - Interactive Elements Answer Key Ionic Bonds Review And Reinfore
- 8. Staying Engaged with Answer Key Ionic Bonds Review And Reinfore
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer Key Ionic Bonds Review And Reinfore
- 9. Balancing eBooks and Physical Books Answer Key Ionic Bonds Review And Reinfore
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer Key Ionic Bonds Review And Reinfore
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answer Key Ionic Bonds Review And Reinfore
 - Setting Reading Goals Answer Key Ionic Bonds Review And Reinfore
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answer Key Ionic Bonds Review And Reinfore
 - Fact-Checking eBook Content of Answer Key Ionic Bonds Review And Reinfore
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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